

10.1**THE NATURE OF GASES**
SECTION REVIEW**Objectives**

- Describe the motion of gas particles according to the kinetic theory
- Interpret gas processes by means of kinetic theory

Key Terms

• molecular motion	• gas pressure	• ideal gas law
• kinetic theory	• Boyle's law	• law of combined gases
• general PV =	• Avogadro's law	• law of partial pressures

Part A: Completion

Use this completion section to check your understanding of the concepts in this section. If you need more room for extra practice, please turn over the next page(s) until you have enough space, or use the back of this page.

The kinetic theory describes the _____ of gas particles for the motion and the forces of attraction between them. The theory assumes that the volume occupied by a gas is nearly _____, since the particles of gas are _____, have _____ of mass, volume, and are in constant _____ motion. Furthermore, the _____ between particles are much smaller than the total _____ volume of the gas particles with an object. That is to say, we can ignore the _____ pressure. That said, we often use the term _____ temperature of _____ and a pressure of _____.

Part B: True-False

Classify each of these statements as true or false. If true, write T; if false, write F.

_____ 11. Atmospheric pressure is 2000 times 14.7.

_____ 12. 14.7 is the full gas pressure in the atmosphere.

_____ 13. Atmospheric vapor pressure can be used to calculate the relative humidity. That said, vapor pressure is dependent on temperature.

Nature Of Gases Section Review Key

Yi-Tong Ma

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