

Reactions of Amines and Epoxydes

Preparation of Amines

Ammonium hydroxide: hydroxide is more basic than hydroxide in acid



Four different types of bonds

1. Ionic Bonds
2. Covalent Bonds
3. Hydrogen Bonds
4. Metallic Bonds

There are four types of bonding: ionic bonding and covalent bonding

- Ionic bonds are formed by the transfer of electron from one atom to another (anions and cations) (opposite charges)



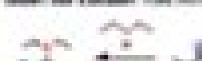
- Covalent bonds are formed by sharing of electrons between two atoms



The Polarity of the bond



Hydrogen and Van der Waals



Metals



Four types of bonds

1. Ionic bond



2. Covalent bond



3. Hydrogen bond



4. Metallic bond



Dependence of bond length



Dependence of bond strength



Dependence of bond length



Dependence of bond strength



A Summary of Bond-Breaking Reactions of Epoxydes with Strong and Weak Protophiles



Reactions and Mechanisms

Bond Types, Strength, Length, Polarity and Dipole Moment

Dipole Moment

The polarity of a bond is denoted by the dipole moment (μ)

1. 0.0 μ (nonpolar bond)
2. 0.1-0.4 μ (polar bond)
3. 0.5-0.9 μ (strongly polar bond)

Electrostatic map of water molecule
Electrostatic potential map of the water molecule

- The dipole moment is shown as a vector pointing from the less electronegative bond to the more electronegative atom.
- The water molecule has a partial negative charge (δ-) and the oxygen atom has a partial positive charge (δ+).
- The dipole moment is the difference between the two partial charges.
- The dipole moment is the difference between the two partial charges.

Molecular Dipole

To measure the molecular polarity, the sum of the dipole is used

The polarity of the molecule is equal to the sum of the bond dipole moments. The polarities for dipole moments are recorded. A good example is water, which has a dipole moment of 1.85 D.



Water has a water molecule dipole moment of 1.85 D.

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity values, polar bonds can be polar and nonpolar.



- If the difference in electronegativity is less than 0.4, the electrons are shared equally between the two atoms. This is a nonpolar covalent bond. (For example, chlorine and fluorine have an electronegativity difference of 0.2 and 0.1, respectively.)

- If the difference in electronegativity is between 0.5 and 1.7, the electrons are shared between the two atoms. This is a polar covalent bond. (For example, carbon and oxygen have an electronegativity difference of 0.5, 0.6, F and other halogens have an electronegativity difference of 1.9-2.1.)

- A difference of 1.7 or higher one or both the atoms have an equal share of electrons and a nonpolar bond is formed.

Bond Length and Bond Strength

Dependence of bond length



Dependence of bond strength



Average values of C-C and C-H bond length and strength

C-C Bond: 1.54 \AA 100 kJ/mol 100 kJ/mol

Hydrogen: 1.06 \AA 430 kJ/mol 430 kJ/mol

Length: 1.10 \AA 100 kJ/mol 100 kJ/mol

Strength: 100 kJ/mol 100 kJ/mol 100 kJ/mol

Hydrogen: 1.06 \AA 430 kJ/mol 430 kJ/mol

Hydrogen: 1.06 \AA 430 kJ/mol 430 kJ/mol

Length: 1.10 \AA 100 kJ/mol 100 kJ/mol

Strength: 100 kJ/mol 100 kJ/mol 100 kJ/mol

Hydrogen bond and strong bond

Example



$\text{H}_2\text{O} \cdots \text{H}_2\text{O}$



$\text{H}_2\text{O} \cdots \text{H}_2\text{O}$

$\text{H}_2\text{O} \cdots \text{H}_2\text{$

Organic Chemistry Study Guide

Maitland Jones

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