

Reactions of Amines and Epoxydes

Preparation of Amines

Ammonium hydroxide: $\text{NH}_4^+ + \text{OH}^- \rightarrow \text{NH}_3 + \text{H}_2\text{O}$



Four different types of bonds



There are four types of bonding: ionic bonding and covalent bonding.

- Ionic bonds are formed by the transfer of electrons from one atom to another (anions and cations) (opposite charges).



Covalent ionic and polar covalent bonds

- Covalent bonds are classified as ionic (polar) and non-polar.
- Ionic bonds are formed by the transfer of electrons.
- Non-polar bonds are formed by the sharing of electrons.
- Polar covalent bonds are formed by the sharing of electrons.
- Non-polar covalent bonds are formed by the sharing of electrons.



Non-polar covalent bonds are stronger than ionic.

Strength of different bonds

Ionic bonds



Covalent bonds



Polar covalent bonds



- If the difference in electronegativity is less than 0.7, the electrons are shared equally between the two atoms. Forming a covalent non-polar bond (like covalent bonds for the C-C and C-H bonds).
- If the difference in electronegativity is between 0.3 and 1.7, one atom has a partial negative charge and the other has a partial positive charge (like ionic bonds for the Na-Cl bond).
- A difference of 1.7 or higher one or both of the atoms have a larger share and a polar bond is formed.

A Summary of Ring-opening Reactions of Epoxydes with Strong and Weak Protophilic Agents

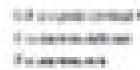


Reactions and Mechanisms

Bond Types, Strength, Length, Polarity and Dipole Moment

Dipole Moment

The polarity of a bond is described by the dipole moment (μ).



Electrostatic potential map of the carbon atom.

- The dipole moment is shown as a vector pointing from the less electronegative atom to the more electronegative atom.
- The vector is proportional to the partial negative charge and the partial positive charge.
- The dipole moment is zero for a polar covalent bond and for a non-polar covalent bond.
- The dipole moment is zero for ionic bonds.

Molecular Dipole

To measure the molecular polarity, the total sum of the dipole is used.

The polarity of the molecule is equal to the sum of the bond dipole moments. If the dipole moments are equal, it is just simple addition.



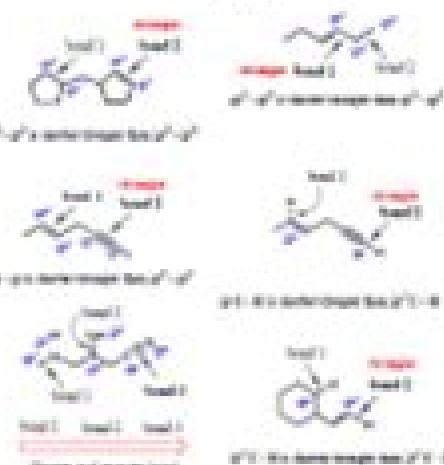
Sum of bond dipole moments for the molecule.

Average values of C-C and C-H bond depending on the molecule

C-C Bond	$\text{H}_2\text{C}=\text{CH}_2$	$\text{CH}_3\text{CH}_2\text{CH}_3$	$\text{CH}_3\text{CH}_2\text{OH}$
Hybridization	sp^2 (polar)	sp^3 (polar)	sp^3 (polar)
Length	1.37 Å	1.54 Å	1.75 Å
Strength	10.0 kJ/mol	17.1 kJ/mol	20.0 kJ/mol
Hybridization	sp^2	sp^3	sp^3
Strength	10.0 kJ/mol	17.1 kJ/mol	20.0 kJ/mol
Length	1.37 Å	1.54 Å	1.75 Å
Strength	10.0 kJ/mol	17.1 kJ/mol	20.0 kJ/mol

Dipole moment and bond length

Example



Bond Length and Bond Strength

Reactions of bond length and strength



Weak protophilic agent ($\text{C}_2\text{H}_5\text{O}$)

Strong protophilic agent (H_2O)

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Strong protophilic agent (H_2O)

Reactions of bond length and strength

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