

Reactions of Amines and Epoxides

Preparation of Amines

Amines are often made by the Gabriel phthalimide synthesis. This involves a series of reactions to form the phthalimide intermediate.



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Epoxide Reactions and Mechanisms

Bond Types, Strength, Length, Polarity and Dipole Moment

Types of Bonding Interactions

There are two types of bonding interaction: covalent bonding and ionic bonding.

- **Covalent bonds** are formed by the sharing of electron pairs between two atoms. They are stronger than ionic bonds (predominant in ionic compounds).



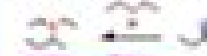
- **Ionic bonds** are formed by the transfer of electron pairs from one atom to another, resulting in the formation of oppositely charged ions.



The Relative Strength of Bonds



Bond and Bonding Interactions



Bonding Interactions



There are two types of bonding interactions:

1. **Covalent bonding**



Sigma and pi bonds

Covalent bonds are classified as sigma (σ) and pi (π) bonds.

- **Sigma bonds** are formed by the end-to-end overlap of atomic orbitals (AOs).
- **Pi bonds** are formed by the side-to-side overlap of atomic orbitals (AOs).
- **Sigma bonds** are stronger than pi bonds.
- **Pi bonds** are weaker than sigma bonds.



• **Sigma bonds** are stronger than pi bonds.

Polar and Nonpolar Covalent Bonds

Depending on the difference in electronegativity, covalent bonds can be polar or nonpolar.



- If the difference in electronegativity is **greater than 1.7**, the bond is **ionic**.
- If the difference in electronegativity is **between 0.5 and 1.7**, the bond is **polar covalent**.
- If the difference in electronegativity is **less than 0.5**, the bond is **nonpolar covalent**.

Examples of polar and nonpolar covalent bonds are given below.

- A difference of **1.7 or higher** is **highly polar** (ionic bond).
- A difference of **0.5 or higher** is **moderately polar** (polar covalent bond).
- A difference of **0.5 or lower** is **nonpolar** (nonpolar covalent bond).

Dipole Moment

The polarity of a bond is described by the **dipole moment** (μ).



- The dipole moment is a vector pointing from the **positive end** to the **negative end** of the bond.
- The **size** of the dipole moment is a measure of the **strength** of the bond.
- The **direction** of the dipole moment is a measure of the **polarity** of the bond.
- **Electronegativity** is a measure of the **ability** of an atom to attract electrons.
- **Electronegativity** is a measure of the **ability** of an atom to attract electrons.

Molecular Dipole

To determine the **molecular dipole**, the **vector sum** of the **bond dipoles** is used.

The **polarity** of the molecule is a measure of the **strength** of the **bond**. The **direction** of the **dipole moment** is a measure of the **polarity** of the **bond**.



• **The molecular dipole moment** is the **vector sum** of the **bond dipoles**.

Bond Length and Bond Strength

• **The length of the bond** is a measure of the **strength** of the **bond**.



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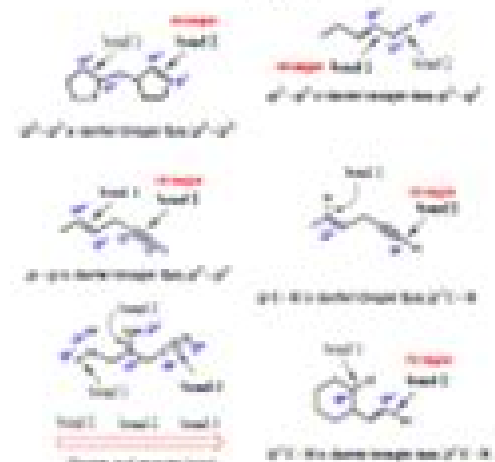


• **The length of the bond** is a measure of the **strength** of the **bond**.

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Length	154 pm	134 pm	121 pm
Strength	347 kJ/mol	460 kJ/mol	591 kJ/mol
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp$
C-H Bond	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Length	109 pm	108 pm	106 pm
Strength	413 kJ/mol	460 kJ/mol	591 kJ/mol

• **The length of the bond** is a measure of the **strength** of the **bond**.

Examples



• **The length of the bond** is a measure of the **strength** of the **bond**.

• **The length of the bond** is a measure of the **strength** of the **bond**.

Organic Chem Amine Study Guide

Steven S. Zumdahl, Paul B. Kelter



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