

Reactions of Amines and Epoxides

Preparation of Amines

Amines are colorless, basic, and have a strong, fishy odor. They are prepared by the reduction of nitriles.



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Bond Types, Strength, Length, Polarity and Dipole Moment

Types of Bonding Interactions

There are two types of bonding: ionic bonding and covalent bonding.

- **Ionic bonds** are formed by the transfer of electrons from one element to another (ionization potential) and electron affinity (electron).



- **Covalent bonds** are formed by sharing of electrons between two atoms.



The Relative Strength of Bonds



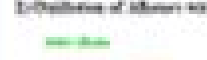
Bond and Bonding Interactions



Bonding Interactions



There are two types of bonding interactions



Sigma and pi bonds

Covalent bonds are classified into sigma and pi bonds.

- **Sigma bonds** are formed by head-on overlap of orbitals.
- **pi bonds** are formed by side-on overlap of orbitals.
- **Single bonds** are sigma bonds.
- **Double bonds** are sigma and pi bonds.



pi bonds are stronger than sigma bonds.

Polar and Nonpolar Covalent Bonds

Depending on the difference in electronegativity, covalent bonds can be polar and nonpolar.



• If the difference in electronegativity is less than 0.5, the electrons are shared equally (nonpolar). For two atoms, forming a nonpolar covalent bond. Two important examples are the C-C and C-H bonds.

• If the difference in electronegativity is between 0.5 and 1.7, we have a polar covalent bond. Most common examples are bonds for carbon to oxygen, N, O, F and other halogens.

• A difference of 1.7 or higher can be high that the electrons are not shared (ionic) and we have an ionic bond.

Dipole Moment

The polarity of a bond is described by the dipole moment (μ).



- The dipole moment is shown as a vector pointing from the less electronegative to the more electronegative atom.

- The more electronegative atom has a partial negative charge (δ⁻) and the less electronegative atom has a partial positive charge (δ⁺).

- **δ⁻** (delta minus) indicates an electron deficient atom.

- **δ⁺** (delta plus) indicates an electron rich atom.

Molecular Dipole

To determine the molecular polarity, the vector sum of the dipoles is used.

The geometry of the molecule is important. Symmetrical molecules are nonpolar even if the dipoles are present. A good example is carbon dioxide.



Cancel the vector as the other pointing point in the same way.

Bond Length and Bond Strength

The longer the bond, the weaker it is.

Bond	H-H	H-Cl	H-Br	H-I
Length	109 pm	127 pm	141 pm	154 pm
Strength	436 kJ/mol	431 kJ/mol	366 kJ/mol	298 kJ/mol

Double bonds are stronger than single bonds. Double bonds are stronger than single bonds.

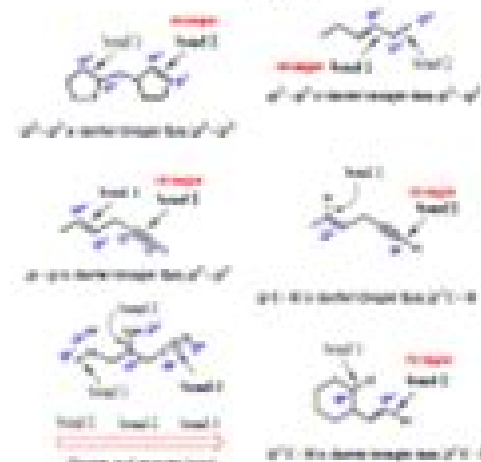


The longer the bond, the weaker it is.

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Length	154 pm	134 pm	121 pm
Strength	347 kJ/mol	461 kJ/mol	591 kJ/mol
Hybridization	sp^3-sp^2	sp^2-sp	$sp-sp$
Length	150 pm	143 pm	128 pm
Strength	356 kJ/mol	479 kJ/mol	591 kJ/mol

Higher s character = stronger and shorter bond

Examples



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