

# Reactions of Amines and Epoxydes

## Preparation of Amines

Ammonium hydroxide:  $\text{NH}_4^+ + \text{OH}^- \rightarrow \text{NH}_3 + \text{H}_2\text{O}$



## Four different types of bonds



There are four types of bonding between atoms in a molecule.

- Ionic bonds are formed by the transfer of electron from one atom to another (anions and cations) (opposite charges).



- Covalent bonds are formed by sharing of electrons between two atoms.



## Region (p) and p-orbital bonds

- Covalent bonds are classified as region (p) and p-orbital.
- Region bonds are formed by bond between two atoms.
- p bonds are formed by side-on overlap of orbitals.
- Very simple bond is a p-orbital.
- Non-overlapping bond consists of two regions and two p-orbitals.



Region bonds are stronger than p-orbital.

## Four types of bonds

## Definition of different bonds

## Hydrogen bonds

## Metallic bonds

## Ionic bonds

## Covalent bonds

## Polar and Nonpolar Covalent Bonds



- If the difference in electronegativity is less than 0.4, the electrons are shared equally between the two atoms. Forming a nonpolar covalent bond (like a C-C bond) or partial charges on the C-C bond.
- If the difference in electronegativity is between 0.5 and 1.7, one atom has a partial negative charge and the other has a partial positive charge (like H-Cl bond).
- A difference of 1.7 or higher one or both the atoms have a larger shared and a lesser bond is formed.

## A Summary of Ring-Opening Reactions of Epoxydes with Strong and Weak Nucleophiles

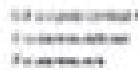


### Reactions and Mechanisms

## Bond Types, Strength, Length, Polarity and Dipole Moment

### Dipole Moment

The polarity of a bond is denoted by the dipole moment ( $\mu$ ).



Direction may be indicated by arrows on the vector sum.

- The dipole moment is shown as a vector pointing from the less electronegative toward the more electronegative atom.
- The vector dipole moment is the partial negative charge and the partial positive charge.
- The dipole plot indicates the charge difference.
- The dipole moment indicates the charge and area.

### Molecular Dipole

To measure the molecular polarity, the total sum of the dipole is used.

The polarity of the molecule is equal to the sum of molecular dipole moments for dipole moment are canceled. A good example is water, dipole.



Water has water as the molecular dipole for the water.

## Bond Length and Bond Strength

### Reactions of bond to length

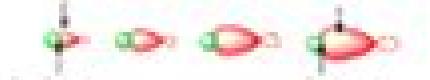


Length:  $\text{Na}^+ - \text{Cl}^- = 2.27 \text{ \AA}$ ,  $\text{C} = \text{O} = 1.21 \text{ \AA}$ ,  $\text{C} - \text{O} = 1.47 \text{ \AA}$ ,  $\text{C} - \text{H} = 1.10 \text{ \AA}$ .

Strength:  $\text{Na}^+ - \text{Cl}^- = 100 \text{ kJ/mol}$ ,  $\text{C} = \text{O} = 870 \text{ kJ/mol}$ ,  $\text{C} - \text{O} = 770 \text{ kJ/mol}$ ,  $\text{C} - \text{H} = 410 \text{ kJ/mol}$ .

Water which has four bonds, three in the water and one oxygen bond.

Water molecule has three region bonds and one p-orbital bond.



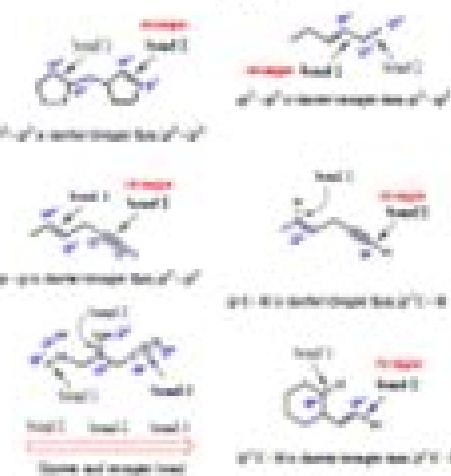
### Reactions of bond to length

## Average values of C-C and C-H bond depending on hybridization

C-C Bond	$\text{sp}_3$ - $\text{sp}_3$	$\text{sp}_2$ - $\text{sp}_2$	$\text{sp}$ - $\text{sp}$
Hybridization	$\text{sp}^3$ - $\text{sp}^3$	$\text{sp}^2$ - $\text{sp}^2$	$\text{sp}$ - $\text{sp}$
Length	1.54 \AA	1.35 \AA	1.10 \AA
Strength	100 kJ/mol	177 kJ/mol	390 kJ/mol
Hybridization	$\text{sp}^3$	$\text{sp}^2$	$\text{sp}$
Strength	100 kJ/mol	177 kJ/mol	390 kJ/mol
Length	1.10 \AA	1.35 \AA	1.54 \AA
Strength	100 kJ/mol	177 kJ/mol	390 kJ/mol

## Water is the water and stronger bond

## Example



# Organic Chem Study Guide

**Marc Loudon**

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