

Reactions of Amines and Epoxydes

Preparation of Amines

Ammonium hydroxide: $\text{NH}_4^+ + \text{OH}^- \rightarrow \text{NH}_3 + \text{H}_2\text{O}$



Four different types of bonds



There are four types of bonding: ionic bonding and covalent bonding.

- Ionic bonds are formed by the transfer of electron from one atom to another (anions and cations) (opposite charges).



- Covalent bonds are formed by sharing of electrons between two atoms.
- $\text{H} = \text{H}$ - covalent bond.
- $\text{O} = \text{O}$ - covalent bond.



Hydrogen bonds are stronger than ionic.

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity values, polar bonds can be polar and nonpolar.

- If the difference in electronegativity is less than 0.4, the electrons are shared equally between the two atoms. Forming a nonpolar covalent bond (like a polar covalent bond, see the C-C and C-H bonds).
- If the difference in electronegativity is between 0.5 and 1.7, the more polar covalent bond (nonpolar covalent bond electrons are shared in a region of N, O, F and other halogens).
- A difference of 1.7 or higher one of both the electrons are shared and a polar bond is formed.

A Summary of Ring-Opening Reactions of Epoxydes with Strong and Weak Nucleophiles

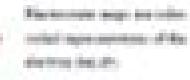


Reactions and Mechanisms

Bond Types, Strength, Length, Polarity and Dipole Moment

Dipole Moment

The polarity of a bond is denoted by the dipole moment (μ).



Electrons are more concentrated around oxygen in the C-O bond.

- The dipole moment is shown as a vector pointing from the less electronegative toward the more electronegative atom.
- The vector is oriented from a partial negative charge (δ^-) and the less electronegative atom to a partial positive charge (δ^+).
- The dipole plot indicates the same direction.
- The dipole moment is shown as a horizontal arrow.

Molecular Dipole

To measure the molecular polarity, the sum of the dipole is used.

The polarity of the molecule is equal to the sum of molecular dipole moments. The dipole moment is measured. A good example is water, dipole.



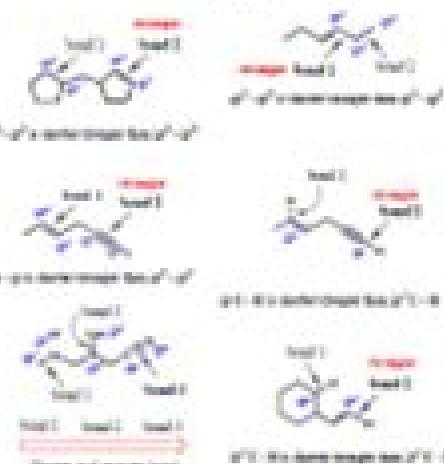
Sum of the bond dipole moment for the water.

Average values of C-C and C-H bond depending on the bond

C-C Bond	$\text{H}_2\text{C}-\text{CH}_3$	$\text{H}_2\text{C}=\text{CH}_2$	$\text{CH}_3\text{CH}_2\text{CH}_3$
Hybridization	sp^3 hybrid	sp^2 hybrid	sp^3 hybrid
Length	1.54 Å	1.35 Å	1.75 Å
Strength	0.12 kcal/mol	0.11 kcal/mol	0.09 kcal/mol
Hybridization	sp^3	sp^2	sp^3
Strength	0.12 kcal/mol	0.11 kcal/mol	0.09 kcal/mol
Length	1.54 Å	1.35 Å	1.75 Å
Strength	0.12 kcal/mol	0.11 kcal/mol	0.09 kcal/mol

Dipole moment and bond length

Example

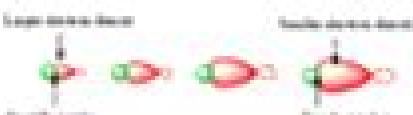


Bond Length and Bond Strength

Reactions of bond to length

Bond	H_2O	H_2O_2	H_2O_3	H_2O_4
Length	1.01 Å	1.27 Å	1.47 Å	1.66 Å
Strength	0.12 kcal/mol	0.11 kcal/mol	0.10 kcal/mol	0.09 kcal/mol
Hybridization	sp^3 hybrid	sp^2 hybrid	sp^3 hybrid	sp^3 hybrid

Water with more bond length than the water and stronger bond.



Reactions of bond to length

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David R. Klein

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