

Reactions of Amines and Epoxides

Preparation of Amines

Amination of alcohols: Polyalkylation is a serious problem for indirect synthesis of amines



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Regioselectivity and Stereochemistry

Bond Types, Strength, Length, Polarity and Dipole Moment

Two types of covalent bond

There are two types of covalent bond: **polar covalent bond** and **non-polar covalent bond**

- Polar covalent bond:** formed by the transfer of electrons from one element to another (electronegativity difference) and **non-polar covalent bond** (identical elements)



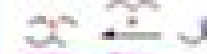
- Covalent bond:** formed by sharing of electrons between two atoms



The difference between sigma and pi bonds



Bond and Electron sharing



Bonding energy



There are two types of covalent bond

Definition of sigma and pi bonds



Sigma and pi bonds

Covalent bonds are classified into **sigma** and **pi** bonds

- Sigma bonds:** are formed by head-on (axial) overlap of orbitals
- Pi bonds:** are formed by side-on (lateral) overlap of orbitals
- Sigma bonds:** are stronger than pi bonds
- Pi bonds:** are weaker than sigma bonds



Pi bonds are weaker than sigma bonds

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity

covalent bonds can be **polar** and **nonpolar**



- If the difference in electronegativity is **less than 0.5**, the electrons are shared equally (nonpolar covalent bond), forming a **nonpolar covalent bond**
- If the difference in electronegativity is **between 0.5 and 1.7**, the electrons are shared unequally (polar covalent bond), forming a **polar covalent bond**
- If the difference in electronegativity is **greater than 1.7**, the electrons are transferred from one atom to another, forming an **ionic bond**

A difference of **1.7 or higher** is so high that the electrons are not longer shared but a **new bond** is formed

Dipole Moment

The polarity of a bond is described by the **dipole moment** (μ)



- The dipole moment is a vector pointing from the **less electronegative** to the **more electronegative** atom
- The vector electronegativity is a **partial positive charge** (δ^+) and the **partial negative charge** (δ^-)
- Electronegativity** indicates an electron's ability to attract electrons
- Electronegativity** indicates an electron's ability to attract electrons

Molecular Dipole

To determine the molecular polarity, the **vector sum** of the dipoles is used

The geometry of the molecule is important. **symmetrical** molecules are **nonpolar** (even for dipole moments are cancelled). A good example is carbon dioxide



Cancel the vector or the other pointing point in the same way

Bond Length and Bond Strength

The longer the bond, the weaker the bond



Double bonds are stronger than single bonds, and triple bonds are stronger than double bonds

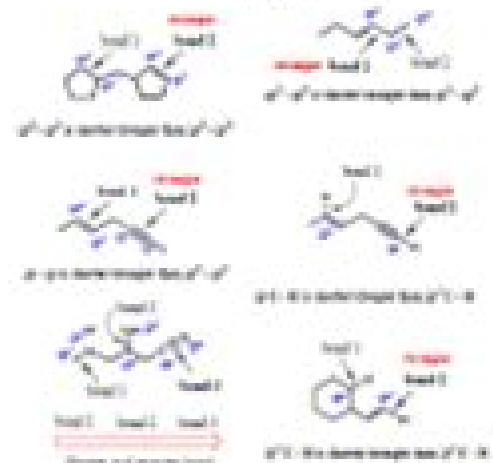


The longer the bond, the weaker the bond

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Length	154 pm	134 pm	121 pm
Strength	347 kJ/mol	461 kJ/mol	591 kJ/mol
Hybridization	sp^3-sp^2	sp^2-sp	$sp-sp$
C-H Bond	109 pm	108 pm	106 pm
Length	154 pm	134 pm	121 pm
Strength	347 kJ/mol	461 kJ/mol	591 kJ/mol

Higher s character = stronger and shorter bond

Examples



Organic Chem Final Study Guide

SA Adler



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