

Name _____ Date _____ Class _____

CHAPTER 12

STUDY GUIDE FOR CONTENT MASTERY

Section 12.3 Limiting Reactants

In your textbook, read about why reactions stop and how to determine the limiting reactants.

Study the diagram showing a chemical reaction and the chemical equation that represents the reaction. Then complete the table. Show your calculations for questions 15–17 in the space below the table.



$$\text{O}_2 = 32\text{ g} = 32\text{ g}$$

The molar masses of Cl_2 , NO_2 , and NO_3 are 71.00 g/mol, 63.01 g/mol, and 83.01 g/mol, respectively.

Amount of Cl_2	Amount of NO_2	Amount of NO_3	Identifying Reactant	Amount and Name of Limiting Reactant
1 molecule	1 molecule	1 molecule	none	none
4 molecules	4 molecules	4 molecules	NO_2	4 molecules NO_2
7 molecules	8 molecules	6 molecules	O_3	6 molecules NO_2
1.00 mol	1.00 mol	1.00 mol	NO_2	1.00 mol NO_2
1.00 mol	0.50 mol	1.00 mol	NO_2	0.50 mol O_3
1.00 mol	0.00 mol	1.00 mol	NO_2	1.00 mol O_3
1.00 mol	0.00 mol	1.00 mol	O_3	1.00 mol NO_2
0.000 mol	0.000 mol	1.00 mol	NO_2	0.000 mol O_3
10.00 g	10.00 g	10.00 g	NO_2	10.00 g NO_2
10.00 g	10.00 g	10.00 g	O_3	10.00 g NO_2
10.00 g	10.00 g	20.76 g	O_3	10.00 g NO_2

balanced equation mole ratio = 2 mol NO_2 /1 mol O_3

$$10.00 \text{ g } \text{O}_3 \times 1 \text{ mol } \text{O}_3/32.00 \text{ g } \text{O}_3 = 0.3125 \text{ mol } \text{O}_3$$

$$20.76 \text{ g } \text{NO}_2 \times 1 \text{ mol } \text{NO}_2/46.01 \text{ g } \text{NO}_2 = 0.4464 \text{ mol } \text{NO}_2$$

actual mole ratio = 0.4464 mol NO_2 /0.3125 mol O_3 = 2.132 mol NO_2 /1.000 mol O_3

Because the actual mole ratio of NO_2 : O_3 is larger than the balanced equation mole ratio of NO_2 : O_3 , there is an excess of NO_2 ; O_3 is the limiting reactant.

Mass of NO_2 used = 0.3125 mol O_3 \times 2 mol NO_2 /1 mol O_3 = 0.6250 mol NO_2

$$0.6250 \text{ mol } \text{NO}_2 \times 46.01 \text{ g } \text{NO}_2/1 \text{ mol } \text{NO}_2 = 28.76 \text{ g } \text{NO}_2$$

Mass of NO_3 produced = 0.6250 mol NO_2 \times 46.01 g NO_2 /1 mol NO_2 = 28.76 g NO_3

Excess NO_2 = 20.00 g NO_2 – 10.00 g NO_2 = 10.00 g NO_2

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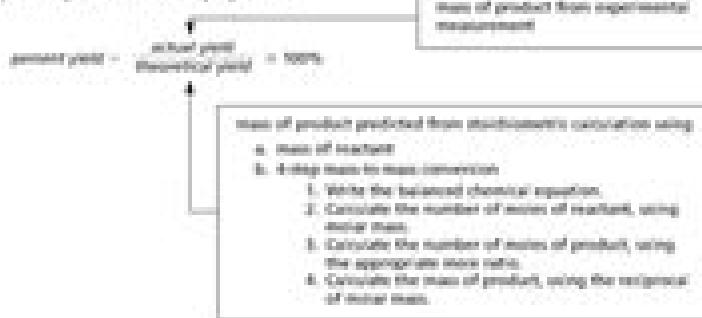
CHAPTER 12

STUDY GUIDE FOR CONTENT MASTERY

Section 12.4 Percent Yield

In your textbook, read about the yields of products.

Study the diagram and the example problem.



Example Problem: The following chemical equation represents the production of palladium, a substance used in the manufacturing of some semiconductor devices.



In one experiment, the student yielded 7.47 g of the oxide from a 20.0-g sample of palladium. Determine the percent yield of this reaction. The molar masses of KlO_3 and Na_2PdCl_4 are 108.73 g/mol and 167.44 g/mol, respectively.

Use the information in the diagram and example problem to evaluate each value or expression below. If the value or expression is correct, write correct. If it is incorrect, write the correct value or expression.

- actual yield: $7.47 \text{ g } \text{K}_2\text{PdCl}_6$
- mass of reactant: 20.0 g KlO_3 correct
- number of moles of reactant: $20.0 \text{ g } \text{KlO}_3 \times \frac{1 \text{ mol } \text{KlO}_3}{108.73 \text{ g } \text{KlO}_3} = 0.183 \text{ mol } \text{KlO}_3$
- number of moles of product: $0.183 \text{ mol } \text{KlO}_3 \times \frac{1 \text{ mol } \text{K}_2\text{PdCl}_6}{1 \text{ mol } \text{KlO}_3} = 0.183 \text{ mol } \text{K}_2\text{PdCl}_6$
- theoretical yield: $0.183 \text{ mol } \text{K}_2\text{PdCl}_6 \times \frac{167.44 \text{ g } \text{K}_2\text{PdCl}_6}{1 \text{ mol } \text{K}_2\text{PdCl}_6} = 30.7 \text{ g } \text{K}_2\text{PdCl}_6$
- percent yield: $\frac{7.47 \text{ g } \text{K}_2\text{PdCl}_6}{30.7 \text{ g } \text{K}_2\text{PdCl}_6} \times 100\% = 24.4\%$

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