

# Reactions of Amines and Epoxides

## Preparation of Amines

Amination of alcohols: Polyalkylation is a serious problem for substituted hydroxyl compounds



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



## Bond Types, Strength, Length, Polarity and Dipole Moment

### Two different bond types

There are two types of bonding: **ionic bonding** and **covalent bonding**

- **Ionic bonds** are formed by the transfer of electrons from one element to another (ionization potential) and electron affinity (electron).



- **Covalent bonds** are formed by sharing of electrons between two atoms.



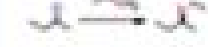
### The Relative Electronegativity of an



### Bonded and Unbonded electron



### Electron movement



### There are two main categories

#### Definition of electron pair



### Sigma (σ) and pi (π) bonds

Covalent bonds are classified in **sigma (σ) bonds** and **pi (π) bonds**

- **Sigma bonds** are formed by head-on (frontal) overlap of orbitals.
- **pi bonds** are formed by side-on overlap of orbitals.
- **One single bond** is a **sigma bond**.
- **Two multiple bonds** consist of one **sigma** and **π** bonds.



π bonds are stronger than σ bonds.

### Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity

values, covalent bonds can be **polar** and **nonpolar**.



- If the difference in electronegativity is **less than 0.5**, the electrons are shared equally (nonpolar bond).
- If the difference in electronegativity is **between 0.5 and 1.7**, we have a **polar covalent bond**. More electronegative atom attracts the electron pair.
- A difference of **1.7 or higher** can be high that the electrons are not longer shared and we have **ionic bond** is formed.

### Dipole Moment

The polarity of a bond is described by the **dipole moment (μ)**

It is a vector quantity



Electronegative atoms are color-coded (negative end of the dipole moment)

- The **dipole moment** is always a vector pointing from the less electronegative to the more electronegative atom.
- The more electronegative atom has a **partial negative charge (δ-)** and the less electronegative atom has a **partial positive charge (δ+)**.
- **δ- (delta minus)** indicates an electron deficient atom.
- **δ+ (delta plus)** indicates an electron rich atom.

### Molecular Dipole

To determine the **molecular polarity**, the vector sum of the **dipole moment**

The geometry of the molecule is important. Symmetrical molecules are nonpolar even if the dipole moments are present. A good example is carbon dioxide.



Consider the vector sum of the dipole moments in the water molecule.

### Bond Length and Bond Strength

As the bond length increases, the bond strength decreases.

Bond	H-H	H-Cl	H-Br	H-I
Length	109 pm	127 pm	141 pm	154 pm
Strength	436 kJ/mol	431 kJ/mol	366 kJ/mol	298 kJ/mol

Double bonds are stronger than single bonds. Triple bonds are stronger than double bonds.

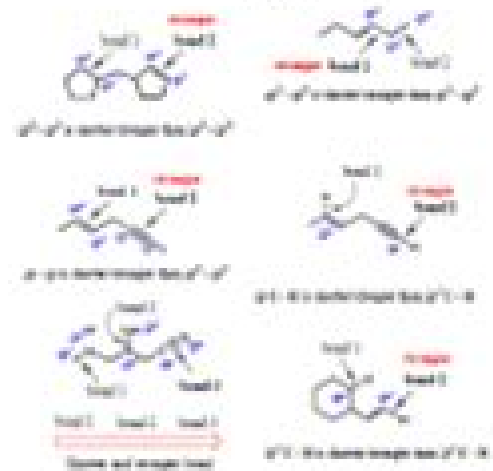


As the bond length increases, the bond strength decreases.

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	$sp^3-sp^3$	$sp^2-sp^2$	$sp-sp$
Hybridization	$sp^3-sp^3$	$sp^2-sp^2$	$sp-sp$
Length	154 pm	134 pm	120 pm
Strength	347 kJ/mol	461 kJ/mol	591 kJ/mol
Hybridization	$sp^3-sp^2$	$sp^2-sp$	$sp-sp$
C-H Bond	109 pm	108 pm	106 pm
Length	154 pm	134 pm	120 pm
Strength	347 kJ/mol	461 kJ/mol	591 kJ/mol

Higher s character = stronger and shorter bond

### Examples



# Organic Chemistry Study Guide Solomon

**Arthur James Wells**



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