

Reactions of Amines and Epoxides

Preparation of Amines

Amines are colorless, basic, and have a strong, fishy odor. They are prepared by the reduction of nitriles.



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Epoxide Reactions and Mechanisms

Bond Types, Strength, Length, Polarity and Dipole Moment

Free-radical halogenation



There are two types of halogen bond: ionic and covalent bond.

• **Ionic bonds** are formed by the transfer of electrons from one element to another (ionization potential) and anion (electron affinity). Examples:



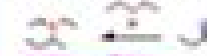
• **Covalent bonds** are formed by sharing of electrons between two atoms.



The Relative Electronegativity of Atoms



Bond and Electron sharing



Bonding electrons



There are two types of halogen bond



Sigma and pi bonds

Covalent bonds are classified into sigma and pi bonds.

• **Sigma bonds** are formed by head-on overlap of orbitals.

• **pi bonds** are formed by side-on overlap of orbitals.

• **Single bond** is a sigma bond.

• **Double bond** consists of one sigma and one pi bond.



pi bonds are stronger than sigma bonds.

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity of atoms, covalent bonds can be polar and nonpolar.

• **Polar covalent bond** is formed when the difference in electronegativity is between 0.5 and 1.7.

• **Nonpolar covalent bond** is formed when the difference in electronegativity is less than 0.5.

• **Electronegativity** is the ability of an atom to attract electrons towards itself. The electronegativity values for various elements are given in the table below.

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Dipole Moment

The polarity of a bond is described by the dipole moment (μ).

It is a vector quantity.



The direction of the dipole moment is from the positive end to the negative end.

• The dipole moment is a vector pointing from the less electronegative atom to the more electronegative atom.

• The more electronegative atom has a partial negative charge (δ^-) and the less electronegative atom has a partial positive charge (δ^+).

• **Electronegativity** is the ability of an atom to attract electrons towards itself.

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Molecular Dipole

To determine the molecular polarity, the vector sum of the dipole moments is calculated.

The polarity of the molecule is determined by the vector sum of the dipole moments of the bonds. A good example is carbon dioxide.



Carbon dioxide is a nonpolar molecule.



Chloromethane is a polar molecule.

Bond Length and Bond Strength

The length of the bond is the distance between the nuclei of the two atoms.

Bond	H-H	H-Cl	H-Br	H-I
Length	109 pm	127 pm	141 pm	154 pm
Strength	436 kJ/mol	431 kJ/mol	366 kJ/mol	298 kJ/mol

The bond length increases as the size of the atoms increases.

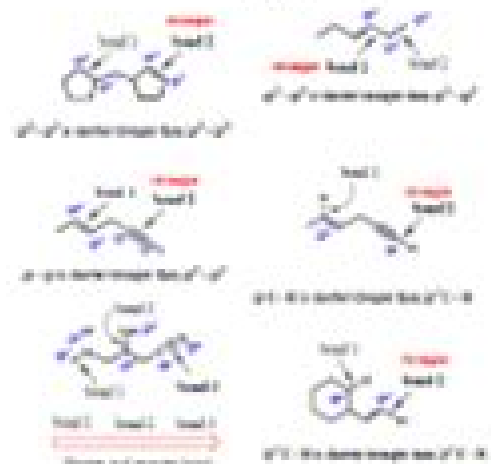


The strength of the bond is the energy required to break the bond.

	Electronegativity	Electronegativity	Electronegativity
C, F Bond	$H_2C=CH_2$	$H_2C=CH_2$	$H_2C=CH_2$
Electronegativity	$\delta^+ - \delta^-$	$\delta^+ - \delta^-$	$\delta^+ - \delta^-$
Length	139 pm	134 pm	133 pm
Strength	445 kJ/mol	472 kJ/mol	485 kJ/mol
Electronegativity	$\delta^+ - \delta^-$	$\delta^+ - \delta^-$	$\delta^+ - \delta^-$
Electronegativity	2.5	3.5	4.0
Length	139 pm	134 pm	133 pm
Strength	445 kJ/mol	472 kJ/mol	485 kJ/mol

Higher electronegativity difference leads to stronger bond

Examples



Chloromethane is a polar molecule.

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