

Reactions of Amines and Epoxides

Preparation of Amines

Amination of alcohols: Polyalkylation is a serious problem for tertiary hydroxyl groups



A Summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Regioselectivity and Stereochemistry

Bond Types, Strength, Length, Polarity and Dipole Moment

Two types of covalent bond

There are two types of covalent bond: bonding and non-bonding bond

- **Bonding bonds** are formed by the overlap of electron orbitals and are classified as **sigma bonds** (single bonds) and **pi bonds** (double bonds).



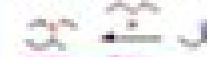
- **Covalent bonds** are formed by sharing of electrons between two atoms.



The Relative Strength of Covalent Bonds



Bonding and Non-bonding Bonds



Bonding and Non-bonding Bonds



There are two types of covalent bond

Definition of sigma bond



Sigma and pi bonds

Covalent bonds are classified as sigma and pi bonds

- **Sigma bonds** are formed by head-on overlap of orbitals.
- **Pi bonds** are formed by side-on overlap of orbitals.
- **Single bonds** are sigma bonds.
- **Double bonds** consist of one sigma and one pi bond.



pi bonds are stronger than sigma bonds

Polar and Nonpolar Covalent Bonds

Depending on the difference in electronegativity

covalent bonds can be polar and nonpolar



- If the difference in electronegativity is **less than 0.5**, the electrons are shared equally (nonpolar covalent bond).
- If the difference in electronegativity is **between 0.5 and 1.7**, the electrons are shared unequally (polar covalent bond).
- If the difference in electronegativity is **greater than 1.7**, the electrons are not shared (ionic bond).

A difference of 1.7 or higher can be high, but the difference can be higher (shared) and a lower bond is formed.

Dipole Moment

The polarity of a bond is described by the **dipole moment** (μ).

It is a vector quantity.



Electronegativity is a relative measure of the ability to attract electrons.

- The dipole moment is a vector pointing from the less electronegative to the more electronegative atom.
- The more electronegative atom has a partial negative charge (δ^-) and the less electronegative atom has a partial positive charge (δ^+).
- **Electronegativity** indicates an electron-attracting atom.
- **Electronegativity** indicates an electron-releasing atom.

Molecular Dipole

To determine the molecular dipole, the vector sum of the bond dipoles is used.

The polarity of the molecule is a vector sum of the bond dipoles. The polarity of the molecule is a vector sum of the bond dipoles.



Consider the vector sum of the bond dipoles to find the molecular dipole.

Bond Length and Bond Strength

The bond length is the distance between the nuclei of the two atoms.

Bond	H-H	H-Cl	H-Br	H-I
Length	109 pm	127 pm	141 pm	154 pm
Strength	436 kJ/mol	431 kJ/mol	366 kJ/mol	298 kJ/mol

Double bonds are stronger than single bonds, and triple bonds are stronger than double bonds.

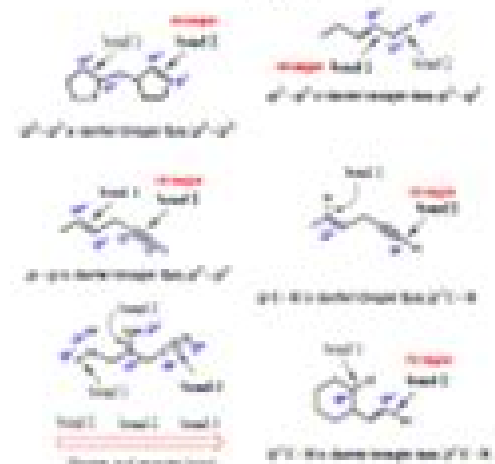


The bond length is the distance between the nuclei of the two atoms.

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Length	154 pm	134 pm	121 pm
Strength	347 kJ/mol	461 kJ/mol	591 kJ/mol
Hybridization	sp^3-sp^2	sp^2-sp	$sp-sp$
Length	147 pm	133 pm	121 pm
Strength	356 kJ/mol	475 kJ/mol	591 kJ/mol

Higher s character means stronger bond

Examples



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