

## 6. $\text{CrO}_3$

### Chromium trioxide

- The mixture of chromium trioxide and sulfuric acid dissolved in a mixture of acetone and water is called **Jones reagent**.
- In aqueous acidic solutions,  $\text{CrO}_3$  forms chromic acid ( $\text{H}_2\text{CrO}_4$ ) which acts, as a strong oxidizing agent.
- It is also used to prepare Pyridinium chlorochromate (**PCC**) (Pyridine +  $\text{CrO}_3$  +  $\text{HCl}$ ) which is a mild oxidizing agent.

Function –

Mild oxidizing agent	Strong oxidizing agent
1. converts primary alcohols to Aldehydes	1. converts primary alcohols to carboxylic acids
2. secondary alcohols to ketones.	2. secondary alcohols to ketones.
3. NO oxidation of tertiary alcohols	3. NO oxidation of tertiary alcohols

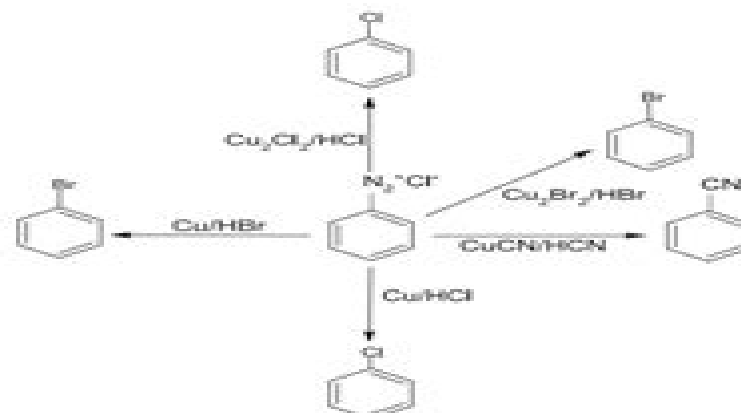


### Reaction of alcohols with $\text{CrO}_3$

1. Oxidation of primary alcohols to aldehydes or carboxylic acids.
2. Oxidation of secondary alcohols to ketones.
3. Oxidation of tertiary alcohols to no reaction.
4. Oxidation of alcohols to aldehydes or ketones.
5. Oxidation of alcohols to aldehydes or ketones.

### Reaction of alcohols with $\text{CrO}_3$

1. Oxidation of primary alcohols to aldehydes or carboxylic acids.



### Sandmeyer's Reaction



### Gatterman's Reaction



2. It is also used for preparing organocuprates (Gilman reagents)



### Example

### Reaction scheme

1. Oxidation of primary alcohols to aldehydes or carboxylic acids.

2. Oxidation of secondary alcohols to ketones.



# Organic Chemistry Reagent Study Guide

**M Tight**



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