

Reactions of Amines and Epoxides

Preparation of Amines

Amines are colorless, basic, and have a strong, fishy odor.



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Regioselectivity and Stereochemistry

Bond Types, Strength, Length, Polarity and Dipole Moment

Types of Bonding Interactions

There are two types of bonding: ionic bonding and covalent bonding.

- **Ionic bonds** are formed by the transfer of electrons from one element to another (ionization potential) and electron affinity (electron).



- **Covalent bonds** are formed by sharing of electrons between two atoms.



The Relative Strength of Ions



Bond and Bonding Interactions



Bonding Interactions



There are two types of bonding interactions



Sigma and pi bonds

Covalent bonds are classified as sigma and pi bonds.

- **Sigma bonds** are formed by head-on overlap of orbitals.
- **pi bonds** are formed by side-on overlap of orbitals.
- **Single bonds** are sigma bonds.
- **Double bonds** are sigma and pi bonds.



pi bonds are stronger than sigma bonds.

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity values, covalent bonds can be polar and nonpolar.



- If the difference in electronegativity is less than 0.5, the electrons are shared equally (nonpolar).
- If the difference in electronegativity is between 0.5 and 1.7, the electrons are shared unequally (polar).
- If the difference in electronegativity is greater than 1.7, the electrons are not shared (ionic).

- A difference of 1.7 or higher can be high for the electrons are not shared (ionic).

- A difference of 1.7 or higher can be high for the electrons are not shared (ionic).

Dipole Moment

The polarity of a bond is described by the dipole moment (μ).



- The dipole moment is a vector pointing from the less electronegative to the more electronegative atom.
- The more electronegative atom has a partial negative charge (δ⁻) and the less electronegative atom has a partial positive charge (δ⁺).
- δ⁺ (delta plus) indicates a partial positive charge.
- δ⁻ (delta minus) indicates a partial negative charge.

Molecular Dipole

To determine the molecular polarity, the vector sum of the dipoles is used.

The geometry of the molecule is important. Symmetrical molecules are nonpolar even if the dipoles are present. A good example is carbon dioxide.



Carbon dioxide is a linear molecule and the dipoles cancel.

Bond Length and Bond Strength

The longer the bond, the weaker it is.



Double bonds are stronger than single bonds.

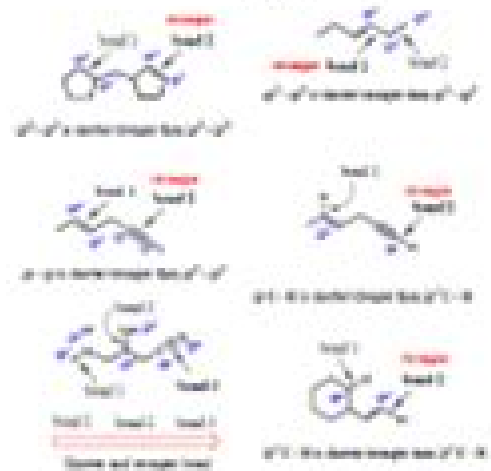


The longer the bond, the weaker it is.

| Average values of C-C and C-H bonds depending on hybridization | | | |
|--|----------------------------------|----------------------------------|------------|
| C-C Bond | sp ³ -C-C | sp ² -C-C | sp-C-C |
| Hybridization | sp ³ -sp ³ | sp ² -sp ² | sp-sp |
| Length | 154 pm | 134 pm | 120 pm |
| Strength | 347 kJ/mol | 461 kJ/mol | 591 kJ/mol |
| Hybridization | sp ³ -sp ² | sp ² -sp | sp-sp |
| Length | 147 pm | 133 pm | 120 pm |
| Strength | 377 kJ/mol | 461 kJ/mol | 591 kJ/mol |

Higher s character means stronger bond

Examples



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