

# Reactions of Amines and Epoxides

## Preparation of Amines

Amines are colorless, basic, and have a strong, fishy odor.



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Regioselectivity and Stereochemistry

## Bond Types, Strength, Length, Polarity and Dipole Moment

### Two types of covalent bonding

There are two types of covalent bonding: **sigma bonding** and **pi bonding**.

- Sigma bonds** are formed by the overlap of atomic orbitals (AOs) in a head-on manner (end-to-end) and are **stronger**.



- Pi bonds** are formed by the overlap of atomic orbitals (AOs) in a side-on manner (side-to-side) and are **weaker**.



### The Relative Strength of $\sigma$ and $\pi$ Bonds



### Bond and Bond Energy



### Bond Energy



### There are two types of covalent bonding

Definition of covalent bond



## Sigma and pi bonds

Covalent bonds are classified into **sigma** and **pi** bonds.

- Sigma bonds** are formed by the head-on overlap of atomic orbitals (AOs).
- Pi bonds** are formed by the side-on overlap of atomic orbitals (AOs).

- Sigma bonds** are stronger than **pi bonds**.
- Sigma bonds** are formed by the head-on overlap of atomic orbitals (AOs).



$\sigma$  bonds are stronger than  $\pi$  bonds.

## Polar and Nonpolar Covalent Bonds

Depending on the difference in electronegativity, covalent bonds can be **polar** or **nonpolar**.

Electronegativity scale (Pauling scale):



- If the difference in electronegativity is **less than 0.5**, the bond is **nonpolar**.
- If the difference in electronegativity is **between 0.5 and 1.7**, the bond is **polar**.
- If the difference in electronegativity is **greater than 1.7**, the bond is **ionic**.

- A difference of **0.5 or higher** is **highly polar**.
- A difference of **1.7 or higher** is **highly ionic**.

## Dipole Moment

The polarity of a bond is described by the **dipole moment** ( $\mu$ ).



- The dipole moment is a vector pointing from the **less electronegative** atom to the **more electronegative** atom.
- The **size** of the dipole moment is a measure of the **charge separation**.
- $\mu = \text{charge} \times \text{distance}$  (in Debye units).
- $\mu = \text{charge} \times \text{distance}$  (in Debye units).

## Molecular Dipole

To determine the **molecular dipole**, the **vector sum** of the **bond dipoles** is used.

The **polarity** of the molecule is a **vector sum** of the **bond dipoles**.



Sum the vectors to find the net dipole moment.

## Bond Length and Bond Strength

The **length** of a bond is a measure of the **distance** between the two atoms.

Bond	H-H	H-Cl	H-Br	H-I
Length	109 pm	127 pm	141 pm	154 pm
Strength	436 kJ/mol	431 kJ/mol	366 kJ/mol	298 kJ/mol

Double bonds are **stronger** than single bonds.

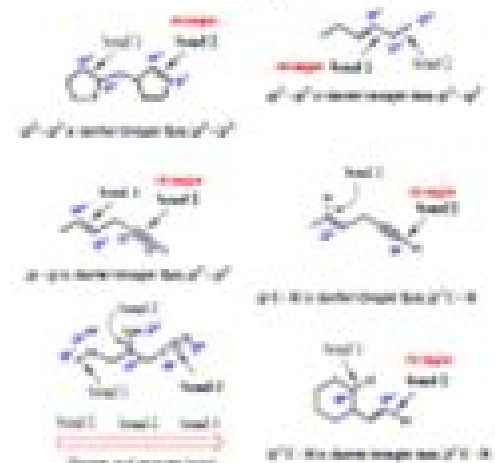


The **length** of a bond is a measure of the **distance** between the two atoms.

	$H_2$	$H_2O$	$H_2S$
<b>Bond Length</b>	109 pm	127 pm	141 pm
<b>Bond Strength</b>	436 kJ/mol	431 kJ/mol	366 kJ/mol
<b>Dipole Moment</b>	0 D	1.85 D	0.95 D
<b>Electronegativity</b>	2.1	3.5	2.5
<b>Length</b>	109 pm	127 pm	141 pm
<b>Strength</b>	436 kJ/mol	431 kJ/mol	366 kJ/mol

Higher electronegativity and shorter bond

## Examples



Stronger and shorter bond

Stronger and shorter bond

Stronger and shorter bond

Stronger and shorter bond

# Organic Chemistry Smith 3rd Study Guide

**John Bingham Roberts**



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## **Table of Contents Organic Chemistry Smith 3rd Study Guide**

1. Understanding the eBook Organic Chemistry Smith 3rd Study Guide
  - The Rise of Digital Reading Organic Chemistry Smith 3rd Study Guide
  - Advantages of eBooks Over Traditional Books
2. Identifying Organic Chemistry Smith 3rd Study Guide
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Organic Chemistry Smith 3rd Study Guide
  - User-Friendly Interface
4. Exploring eBook Recommendations from Organic Chemistry Smith 3rd Study Guide
  - Personalized Recommendations
  - Organic Chemistry Smith 3rd Study Guide User Reviews and Ratings
  - Organic Chemistry Smith 3rd Study Guide and Bestseller Lists

5. Accessing Organic Chemistry Smith 3rd Study Guide Free and Paid eBooks
  - Organic Chemistry Smith 3rd Study Guide Public Domain eBooks
  - Organic Chemistry Smith 3rd Study Guide eBook Subscription Services
  - Organic Chemistry Smith 3rd Study Guide Budget-Friendly Options
6. Navigating Organic Chemistry Smith 3rd Study Guide eBook Formats
  - ePub, PDF, MOBI, and More
  - Organic Chemistry Smith 3rd Study Guide Compatibility with Devices
  - Organic Chemistry Smith 3rd Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Organic Chemistry Smith 3rd Study Guide
  - Highlighting and Note-Taking Organic Chemistry Smith 3rd Study Guide
  - Interactive Elements Organic Chemistry Smith 3rd Study Guide
8. Staying Engaged with Organic Chemistry Smith 3rd Study Guide
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Organic Chemistry Smith 3rd Study Guide
9. Balancing eBooks and Physical Books Organic Chemistry Smith 3rd Study Guide
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Organic Chemistry Smith 3rd Study Guide
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Organic Chemistry Smith 3rd Study Guide
  - Setting Reading Goals Organic Chemistry Smith 3rd Study Guide
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Organic Chemistry Smith 3rd Study Guide
  - Fact-Checking eBook Content of Organic Chemistry Smith 3rd Study Guide
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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