

Reactions of Amines and Epoxydes

Preparation of Amines

Ammonium hydroxide: $\text{NH}_4^+ + \text{OH}^- \rightarrow \text{NH}_3 + \text{H}_2\text{O}$



Four different types of bonds

1. Ionic bond
2. Covalent bond
3. Hydrogen bond
4. Metallic bond

There are four types of bonding: ionic bonding and covalent bonding.

- Ionic bonds are formed by the transfer of electron from one atom to another (anions and cations). (Opposite charges).

Covalent bonds are formed by sharing of electrons between two atoms.

$\text{H} + \text{H} \rightarrow \text{H}_2$ (Ionic bond)

$\text{O}=\text{C}=\text{O}$ (Covalent bond)

Hydrogen bond and pi (p) bonds

- Hydrogen bonds are formed by partial ionic bonding of atoms.
- Covalent bonds are formed by partial sharing of atoms.
- Very simple bond is a covalent bond.
- Very complex bond is a pi (p) bond.

Hydrogen bond
Hydrogen bond is stronger than ionic bond.

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity values, covalent bonds can be polar and nonpolar.

If the difference in electronegativity is less than 0.4, the electrons are shared equally between the two atoms. This is a nonpolar covalent bond. (For example: C-C and C-H bonds).

If the difference in electronegativity is between 0.5 and 1.7, the more polar covalent bond. (Nonpolar covalent bonds are often the carbon-hydrogen in N, O, F and other halogens).

If a difference of 1.7 or higher one or both of the atoms have a larger shared and a polar bond is formed.

A Summary of Ring-Opening Reactions of Epoxydes with Strong and Weak Protagonists



Reactions and mechanisms

Bond Types, Strength, Length, Polarity and Dipole Moment

Dipole Moment

The polarity of a bond is denoted by the dipole moment (μ).

1. 100% ionic bond
2. 50% ionic bond
3. Covalent bond



Electrons are concentrated around the more electronegative atom.

- The dipole moment is shown as a vector pointing from the less electronegative atom to the more electronegative atom.
- The vector is strongest when the partial negative charge and the partial positive charge are at the same point.
- For 100% ionic, electrons are shared 0%.
- For 50% ionic, electrons are shared 50%.

Molecular Dipole

To measure the molecular polarity, the sum of the dipole is used.

The polarity of the molecule is equal to the sum of the bond dipole moments. If the bond dipole moments are equal, it is just simple addition.



Sum of bond dipole moment for the molecule.

Bond Length and Bond Strength

Reactions of bond length and strength



Bond 1: 1.17 \AA Bond 2: 1.27 \AA Bond 3: 1.47 \AA Bond 4: 1.66 \AA

Strength: 10% bond 20% bond 50% bond 100% bond

Length: 1.66 \(\text{\AA}\) 1.47 \(\text{\AA}\) 1.27 \(\text{\AA}\)

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Reactions of bond length and strength

Average values of C-C and C-H bond depending on the molecule

C-C bond: 1.54 \AA 1.54 \AA 1.54 \AA

Hydrogen bond: 1.17 \AA 1.27 \AA 1.47 \AA

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