

9.1. INTRODUCTION

- Potentiometric titration embrace two major types of analyses
- It involves the direct measurement of an electrode potential from which the activity (or concentration) of an active ion may be derived.
- The second type involves measuring the changes in electromotive force (EMF) brought about by the addition of a titrant to the sample.

9.1.1. PRINCIPLE

- The principle involved in the potentiometry is when the pair of electrodes is placed in the sample solution it shows the potential difference by the addition of the titrant or by the change in the concentration of the ions. This can be seen below:

1. Reference electrode/salt bridge/analyte solution/indicator electrode Potentiometer

- The reference electrode is the electrode which consists of its own potential value and it is stable when dipped into the sample solution.
- The salt bridge is used to prevent the interference of the analyte solution with the reference solution.
- Analyte solution is the sample solution whose potential is to be measured.

2. The indicator electrode is the electrode which responds to the changes in the potential of the analyte solution.

The electromotive force of the complete cell is given by the following equation:

$$E_{\text{cell}} = E_{\text{reference}} + E_{\text{indicator}} + E_{\text{junction}}$$

Where $E_{\text{reference}}$ is the electromotive force of the reference electrode;

$E_{\text{indicator}}$ is the electromotive force of the indicator electrode; E_{junction} is the electromotive force at the junction of the liquid.

- The units for the potential are mille volts denoted by mV.

THEORY

- The main theory involved in the potentiometry is when the known potential electrode immersed in the sample solution then the potential is given by the Nernst equation:

$$E = E^{\circ} + (0.592/n) \log c$$

where

- E is the potential of the solution;
- E° is the standard electrode potential;
- n is the valency of the ions;
- c is the concentration of the sample solution;

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Clinical Chemistry Self-assessment Marge A. Brewster,Charles P. Turley,Alan H. B. Wu,1989 **Chemical and Biological Microsensors** Jacques Fouletier,Pierre Fabry,2010-01-26 This book reviews the state of art in the field of chemical sensors for analyses of ionic or molecular species dissolved in liquid media mainly in aqueous solutions The transduction of such devices is based on chemical biological and physical phenomena The fundamental phenomena involved in these sensors are described in the different chapters by specialists having a good expertise in the field Numerous recent bibliographic references are given Most of the devices could be miniaturised using modern technologies allowing a fabrication on a large scale for a mass

production at low cost Moreover such devices could open the field of applications in a near future environmental biomedical food industries domotic and automotive applications etc *World Index of Scientific Translations and List of Translations Notified to ETC.* ,1976 Covers translations of scientific and technical interest from non Western languages into Western languages Medical Laboratory Technology Examination Review Zandra E. Strother,1985 Subject Guide to Books in Print ,1984 Books in Print Supplement ,1988

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Serjeant,1984-06-29 Cells electrodes and basic procedures rudiments of potentiometry classifications of potentiometry and titrimetry classifications and descriptions of electrodes procedures of analytical potentiometry electrodes for potentiotitrimetry Applications to the determination of solution equilibrium data the determination of ionization constants in aqueous solution acid base equilibria in nonaqueous solvents the determination of stability constants by potentiometric titration Applications to titrimetric analysis the basics of titrimetry acid base titrimetry complexometric titrations oxidation reduction titrations precipitation titrations *Advanced Potentiometry* Erzsébet Néher-Neumann,2010-03-14 Advanced Potentiometry details the accurate calculation of potentiometric titrations Additional terms such as the ideal diffusion potential together with the contribution of activity factors furnish calculated stability constants which do not include systematic errors and thus present a complete and correct description of equilibrium processes Moreover serious systematic errors resulting from the recent use of curve fitting computer programs are highlighted later in the text Advanced Potentiometry can advantageously be used by lecturers researchers PhD graduates and undergraduate students working in the fields of Solution Chemistry Physical Chemistry Inorganic Chemistry and Analytical Chemistry and Environmental Research

Improved Equipment for Potentiometric Titration Procedure Edward N. Wise,1950 Class 11-12 Chemistry

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