

Reactions of Amines and Epoxides

Preparation of Amines

Amines are colorless, basic, and have a strong, fishy odor. They are prepared by the reduction of nitriles.



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Bond Types, Strength, Length, Polarity and Dipole Moment

Free-radical halogenation



There are two types of bonding: ionic bonding and covalent bonding.

- **Ionic bonds** are formed by the transfer of electrons from one element to another (ionization potential) and electron affinity (electron).



- **Covalent bonds** are formed by sharing of electrons between two atoms.



The Relative Electronegativity of Atoms



Bond and Resonance Structures



Resonance structures



There are two types of electronegativity

Electronegativity of elements can be measured by the difference in electronegativity.



Sigma and pi bonds

Covalent bonds are classified into sigma and pi bonds.

- **Sigma bonds** are formed by head-on overlap of orbitals.
- **pi bonds** are formed by side-on overlap of orbitals.
- **One single bond is a sigma bond.**
- **Two double bonds consist of one sigma and one pi bond.**



pi bonds are stronger than sigma bonds.

Polar and Nonpolar Covalent Bonds

Depending on the difference in electronegativity, covalent bonds can be polar and nonpolar.



- If the difference in electronegativity is less than 0.5, the electrons are shared equally (nonpolar).
- If the difference in electronegativity is between 0.5 and 1.7, the electrons are shared unequally (polar).
- If the difference in electronegativity is greater than 1.7, the electrons are not shared (ionic).

- A difference of 1.7 or higher can be high for the electrons are not shared (ionic).

- A difference of 1.7 or higher can be high for the electrons are not shared (ionic).

Dipole Moment

The polarity of a bond is described by the dipole moment (μ).



- The dipole moment is a vector pointing from the less electronegative to the more electronegative atom.
- The more electronegative atom has a partial negative charge (δ⁻) and the less electronegative atom has a partial positive charge (δ⁺).
- **δ⁻ (delta minus)** indicates an electron deficient atom.
- **δ⁺ (delta plus)** indicates an electron rich atom.

Molecular Dipole

To determine the molecular polarity, the vector sum of the dipoles is used.

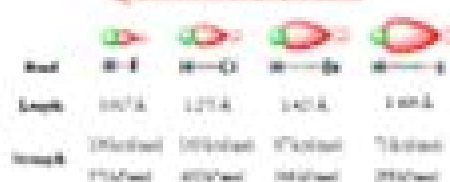
The geometry of the molecule is important. Symmetrical molecules are nonpolar even if the dipoles are present. A good example is carbon dioxide.



Carbon dioxide is a linear molecule and the dipoles cancel.

Bond Length and Bond Strength

The length of a bond is the distance between the nuclei of the two atoms.

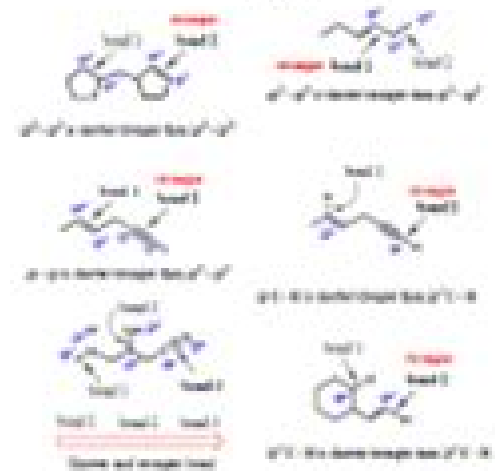


The length of a bond is the distance between the nuclei of the two atoms.

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	sp ³ -C	sp ² -C	sp-C
Hybridization	sp ³ -sp ³	sp ² -sp ²	sp-sp
Length	1.54 Å	1.34 Å	1.20 Å
Strength	33 kcal/mol	112 kcal/mol	160 kcal/mol
Hybridization	sp ³ -sp ³	sp ² -sp ²	sp-sp
Length	1.54 Å	1.34 Å	1.20 Å
Strength	33 kcal/mol	112 kcal/mol	160 kcal/mol

Higher s character = stronger and shorter bond

Examples



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