

# Reading Comprehension

## Compare and Contrast

### Plus Opinion Writing #1

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#### Lions

Lions like to live in open plains or in the savanna of Africa and India. The lion is known as the 'King of Beasts.' Male lions have tan colored coats with dark manes. Lions are good tree climbers, and they can climb high up in a tree. A lioness will have two cubs at a time. She takes care of her cubs for about one year. Lions live and hunt in groups of 10 to 20. Together, they hunt big animals like wildebeest or bison.

There are now fewer lions in the wild, and they are now at risk of becoming extinct.



#### Tigers

Tigers are found only in dense forests or in jungles.

Tigers are orange with brown stripes across their bodies, and they have white fur around the face.

They like to swim, and they are good jumpers. A tigress usually has two to four cubs at a time, and she feeds them until they are about two years old.

Tigers are very powerful and aggressive. They have long teeth, strong jaws, and big paws which allow them to hunt big animals alone.

There are very few tigers left today. They are already endangered, or at risk of becoming extinct.



# Openai Reading Comprehension Compare

**Camilla Rothe**

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web atoms of the same element that differ in their numbers of neutrons are called isotopes many isotopes occur naturally usually one or two isotopes of an element are the most stable and common different isotopes of an element generally have the same physical and chemical properties because they have the same numbers of protons and electrons

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web jan 5 2016 explanation isotopes are atoms of the same element same atomic number ie same number of protons in the nucleus but different atomic masses different mass numbers due to different numbers of neutrons in the nucleus for example we may get an isotope of chlorine with 17 protons and 18 neutrons hence atomic mass 35

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web oct 31 2023 an isotope is one of two or more species of atoms of a chemical element with the same atomic number and position in the periodic table and nearly identical chemical behavior but with different atomic masses and physical properties every chemical element has one or more isotopes

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web verified by topper correct option is d atoms having same atomic number but different atomic mass are known as isotopes the difference in the mass is due to the difference in their number of neutrons with the same number of electrons and protons

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web part i answer the questions based on the above reading what is an isotope isotopes are versions of the same element they have the same number of protons and electrons as the element but different mass numbers and number of neutrons what does the number next to isotopes signify the number indicates the isotope's mass number

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d more than one response is correct e they have different numbers of protons in the nucleus

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web the rate determining slower step gives a rate law showing second order dependence on the  $[A]^2$  concentration and the sum of the two elementary equations gives the overall net reaction in general when the rate determining slower step is the first step in the reaction mechanism the rate law for the overall reaction is the same as the rate

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web the rate determining step limits the overall rate and therefore determines the rate law for the overall reaction introduction multistep reactions one of the most important applications of kinetics is to the study of reaction mechanisms or the sequences of steps by which chemical reactions occur

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web the initial rate law is then measured for each of the reactions consider the reaction between nitrogen monoxide gas and hydrogen gas to form nitrogen gas and water vapor the following data were collected for this reaction at 1280 text o text c see table below notice that the starting concentrations of ce no and ce h

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web determination of a rate law and activation energy in this experiment glass beakers pipettes and a hotplate were utilized to set up the procedures a solution of deionized water a buffer potassium iodide ki sodium thiosulfate na<sub>2</sub>s<sub>2</sub>o<sub>3</sub> starch and hydrogen peroxide h<sub>2</sub>o<sub>2</sub> were used in order to calculate the rate of reaction

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