

1. The diagram below represents the nucleus of an atom.

Key
● = proton
○ = neutron



What are the atomic number and mass number of this atom?

- A) The atomic number is 9 and the mass number is 19.
- B) The atomic number is 9 and the mass number is 20.
- C) The atomic number is 11 and the mass number is 19.
- D) The atomic number is 11 and the mass number is 20.

2. Which statement concerning elements is true?

- A) Different elements must have different numbers of isotopes.
- B) Different elements must have different numbers of neutrons.
- C) All atoms of a given element must have the same mass number.
- D) All atoms of a given element must have the same atomic number.

3. Which subatomic particles are located in the nucleus of an He-4 atom?

- A) electrons and neutrons
- B) electrons and protons
- C) neutrons and protons
- D) neutrons, protons, and electrons

4. Which part of a helium atom is positively charged?

A) electron	B) neutron
C) nucleus	D) orbital

5. Which total mass is the *smallest*?

- A) the mass of 2 electrons
- B) the mass of 2 neutrons
- C) the mass of 1 electron plus the mass of 1 proton
- D) the mass of 1 neutron plus the mass of 1 electron

6. Subatomic particles can usually pass undeflected through an atom because the volume of an atom is composed of

- A) an uncharged nucleus
- B) largely empty space
- C) neutrons
- D) protons

7. An experiment in which alpha particles were used to bombard thin sheets of gold foil led to the conclusion that an atom is composed mostly of

- A) empty space and has a small, negatively charged nucleus
- B) empty space and has a small, positively charged nucleus
- C) a large, dense, positively charged nucleus
- D) a large, dense, negatively charged nucleus

8. Which conclusion is based on the "gold foil experiment" and the resulting model of the atom?

- A) An atom is mainly empty space, and the nucleus has a positive charge.
- B) An atom is mainly empty space, and the nucleus has a negative charge.
- C) An atom has hardly any empty space, and the nucleus has a positive charge.
- D) An atom has hardly any empty space, and the nucleus has a negative charge.

9. Which statement about the mass of an electron is correct?

- A) The mass of an electron is equal to the mass of a proton.
- B) The mass of an electron is less than the mass of a proton.
- C) The mass of an electron is equal to the mass of a neutron.
- D) The mass of an electron is greater than the mass of a neutron.

10. Experiments with gold foil indicated that atoms

- A) usually have a uniform distribution of positive charges
- B) usually have a uniform distribution of negative charges
- C) contain a positively charged, dense center
- D) contain a negatively charged, dense center

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