

13.2**THE NATURE OF LIQUIDS****Section Review****Objectives**

- Identify factors that determine physical properties of a liquid
- Define evaporation in terms of kinetic energy
- Describe the equilibrium between a liquid and its vapor
- Identify the conditions under which boiling occurs

Vocabulary

• vaporization	• boiling point
• evaporation	• normal boiling point
• vapor pressure	

Part A. Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

Liquids are much 1 than gases. Liquids and solids are known as 2 states of matter. The conversion of a liquid to a gas or vapor is called 3. When a liquid that is not 4 changes to a gas, the process is called evaporation. A liquid evaporates faster when heated; however, evaporation itself is a 5 process. When a partially filled container of liquid is sealed, some of the particles at the 6 of the liquid vaporize. These particles collide with the walls of the container, producing a force called 7. The vapor pressure of a liquid can be determined by a device called a 8. A liquid boils when its 9 equals the external pressure. The normal boiling point of a liquid is the temperature at which the vapor pressure is equal to 10.

1. _____
2. _____
3. _____
4. _____
5. _____
6. _____
7. _____
8. _____
9. _____
10. _____

Part B. True-False

Classify each of these statements as always true, AT; sometimes true, ST; or never true, NT.

_____ 11. Particles in a liquid don't have enough kinetic energy to overcome the attractive forces between them and vaporize.

Nature Of Liquids Section Review Key

Sabine Zange

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