

**10.1****THE NATURE OF GASES****SECTION REVIEW****Objectives**

- Describe the motion of gas particles according to the kinetic theory
- Interpret gas pressures in terms of kinetic theory

**Key Terms**

- kinetic energy
- kinetic theory
- gas pressure
- vacuum
- atmospheric pressure
- barometric pressure
- perfect (P<sub>0</sub>)
- standard atmospheric (atm)

**Part A Completion**

Use this completion exercise to check your understanding of the concepts and terms that are highlighted in this section. Write clearly over the completed work a term, short phrase, or sentence.

- The kinetic theory describes the \_\_\_\_\_ of particles in matter and the forces of attraction between them. The theory assumes that the volume occupied by a gas is mostly \_\_\_\_\_, that the particles of gas are \_\_\_\_\_, move \_\_\_\_\_ of each other, and are in constant \_\_\_\_\_ motion. Furthermore, the \_\_\_\_\_ between particles are elastic so that the total \_\_\_\_\_ remains constant. The pressure of a gas results from the countless collisions of the gas particles with an object. Barometers are used to measure \_\_\_\_\_ pressure. Standard conditions are defined as a temperature of \_\_\_\_\_ and a pressure of \_\_\_\_\_.
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**Part B True-False**

Classify each of these statements as true (always true), AT: sometimes true, ST: sometimes true, NT: never true.

- \_\_\_\_\_ 11. Atmospheric pressure is 760 mm Hg.
- \_\_\_\_\_ 12. The SI unit of pressure is the pascal.
- \_\_\_\_\_ 13. Atmospheric pressure decreases as you climb a mountain because the air layer that surrounds the earth thins out.

# Nature Of Gases Section Review Key

**F Rizvi**



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