

## CHAPTER

## STUDY GUIDE

**The Mole****Section 10.1 Measuring Matter***In your textbook, read about counting particles.**In Column B, rank the quantities from Column A from smallest to largest.***Column A**

- 0.5 mol
- 200
- 5
- 6,000,000,000
- $6.02 \times 10^{23}$
- dozen
- four moles
- gross
- pair
- ream

**Column B**

- 1. \_\_\_\_\_
- 2. \_\_\_\_\_
- 3. \_\_\_\_\_
- 4. \_\_\_\_\_
- 5. \_\_\_\_\_
- 6. \_\_\_\_\_
- 7. \_\_\_\_\_
- 8. \_\_\_\_\_
- 9. \_\_\_\_\_
- 10. \_\_\_\_\_

*In your textbook, read about converting moles to particles and particles to moles.**In the boxes provided, write the conversion factor that correctly completes each problem.*11.  
atoms $1.20 \text{ mol Cu} \times \text{ } = 7.22 \times 10^{23} \text{ Cu}$ 12.  
 $\times 10^{-3}$  mol CH<sub>4</sub> $9.25 \times 10^{22} \text{ molecules CH}_4 \times \text{ } = 1.54$ 13.  
mol Xe $1.54 \times 10^{22} \text{ atoms Xe} \times \text{ } = 2.56 \times 10^{-3}$ 14.  
molecules F<sub>2</sub> $3.01 \text{ mol F}_2 \times \text{ } = 1.81 \times 10^{24}$

# Mcgraw Hill Study Guide Answer

**S Baum**

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