

THERMODYNAMICS

review

STUFF TO KNOW

specific heat capacity

enthalpy

entropy

hess's law

gibbs' free law

negative/positive

SPECIFIC HEAT CAPACITY:

$$C_p = q/m \cdot T$$

amount of energy required
to raise the temp of 1g of a
substance by 1 C degree/1K

ENTHALPY

enthalpy change - amount of
energy absorbed by a system
as heat during a process at a
constant pressure (H)

symbol = H

equation: $H =$
 $\Sigma H_{\text{products}} -$
 $\Sigma H_{\text{reactants}}$

ENDOTHERMIC:

(of a reaction or process)
accompanied by or requiring
the absorption of heat
(positive enthalpy)

EXOTHERMIC:

(of a reaction or process)
accompanied by the
release of heat
(negative enthalpy)

HESS' LAW

Used to calculate
enthalpy

symbol = H

equation: $H = H_{\text{products}}$
 $- H_{\text{reactants}}$

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Ishwar K. Puri



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