

Nuclear Chemistry - Guided Practice Questions

Multiple Choice - Choose the one best answer and write its letter in the blank provided.

1) In balancing the nuclear reaction $^{238}_{\text{U}}\text{U} \rightarrow ^{206}_{\text{Pb}}\text{Pb} + \text{?}$, the identity of the missing element is $^{4}_{\text{He}}$. T

(A) Pu. (B) Cl. (C) Th. (D) He. (E) Pb.

2) The following reaction, $^{239}_{\text{Np}}\text{Pu} \rightarrow ^{235}_{\text{U}}\text{U} + \text{?}$, is an example of what form of radioactive decay?

(A) alpha. (B) positron. (C) gamma. (D) beta. (E) $\bar{\nu}$ -capture.

3) The missing product from this reaction is: $^{235}_{\text{U}}\text{U} \rightarrow ^{141}_{\text{Ba}}\text{Ba} + \text{?}$

(A) $^{106}_{\text{Ru}}$ (B) $^{144}_{\text{La}}$ (C) $^{144}_{\text{Ba}}$ (D) $^{146}_{\text{La}}$ (E) $^{146}_{\text{Ru}}$

4) What is the mass number of a neutron _____?

(A) 1. (B) 3. (C) 2. (D) 0. (E) 4.

5) In the nuclear decay reaction represented by $^{14}_{\text{N}}\text{N} \rightarrow ^{14}_{\text{N}}\text{N}$, the emitted particle is _____.

(A) beta particle. (B) neutron. (C) alpha particle. (D) positron. (E) proton.

6) The reaction shown below is responsible for creating C-14 in the atmosphere. What is the bombarding particle _____?



(A) electron. (B) proton. (C) positron. (D) alpha particle. (E) neutron.

7) $^{130}_{\text{I}}\text{I}$ has a half-life of 8.04 days. Assuming you start with a 2.50 mg sample of $^{130}_{\text{I}}$, how many mg will remain after 13.0 days _____?

(A) 0.440. (B) 0.422. (C) 0.298. (D) 0.193. (E) 0.499.

8) All atoms of a given element have the same:

(A) atomic mass. (B) mass number. (C) # of nucleons. (D) atomic #. (E) # of neutrons.

9) Atoms containing radioactive nuclei are called:

(A) radioisotopes. (B) radionuclides. (C) nuclides. (D) nucleons. (E) radioisophones.

10) In what happens to the mass number and the atomic number of an element when it undergoes beta decay?

(A) The mass number decreases by 4 and the atomic number decreases by 2. (B) The mass number does not change and the atomic number decreases by 2. (C) Neither the mass number nor the atomic number change. (D) The mass number increases by 2 and the atomic number increases by 1. (E) The mass number does not change and the atomic number increases by 1.

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