

Chemistry Practice: Balancing Chemical Equations

1. Balance the equations below. Remember this means only using coefficients. All formulas are written correctly.

a.) $\cancel{2} \text{NaHCO}_3 \longrightarrow \text{Na}_2\text{CO}_3 + \text{H}_2\text{O} + \text{CO}_2$

b.) $\text{Zn} + \text{S} \longrightarrow \text{ZnS}$

c.) $\cancel{2} \text{H}_2 + \text{O}_2 \longrightarrow \cancel{2} \text{H}_2\text{O}$

d.) $\cancel{2} \text{FeO} + \text{C} \longrightarrow \cancel{2} \text{Fe} + \text{CO}_2$

e.) $\text{Cl}_2 + \cancel{2} \text{KI} \longrightarrow \cancel{2} \text{KCl} + \text{I}_2$

f.) $\text{H}_2\text{O}_2 + \cancel{2} \text{Mg} \longrightarrow \cancel{2} \text{MgO} + \cancel{2} \text{H}_2$

g.) $\text{Cu} + \cancel{2} \text{H}_2\text{O} \longrightarrow \text{Cu}(\text{OH})_2 + \cancel{2} \text{H}_2$

h.) $\text{Na}_2\text{H}_2\text{O}_2 + \text{H}_2\text{SO}_4 \longrightarrow \cancel{2} \text{H}_2\text{O}_2 + \text{Na}_2\text{SO}_4$

i.) $\text{CuCO}_3 + \cancel{2} \text{H}_2\text{O} \longrightarrow \text{Cu}(\text{OH})_2 + \text{CO}_2$

j.) $\text{NH}_4\text{NO}_3 \longrightarrow \text{N}_2\text{O} + \text{H}_2\text{O}$

k.) $\text{H}_2\text{S} + \text{O}_2 \longrightarrow \text{SO}_2 + \text{H}_2\text{O}$

l.) $\text{H}_2\text{SO}_4 + \cancel{2} \text{NaCl} \longrightarrow \cancel{2} \text{HCl} + \text{Na}_2\text{SO}_4$

m.) $\cancel{2} \text{NaCl} + \cancel{2} \text{H}_2\text{O} \longrightarrow \cancel{2} \text{NaOH} + \text{Cl}_2 + \text{H}_2$

n.) $\cancel{2} \text{BaO} + \cancel{2} \text{Al} \longrightarrow \cancel{2} \text{Ba} + \text{Al}_2\text{O}_3$

o.) $\cancel{2} \text{O}_2 \longrightarrow \cancel{2} \text{O}_2$

p.) $\text{SiO}_2 + \cancel{2} \text{C} \longrightarrow \cancel{2} \text{CO} + \text{Si}$

q.) $\cancel{2} \text{PbS} + \cancel{2} \text{O}_2 \longrightarrow \cancel{2} \text{PbO} + \cancel{2} \text{SO}_2$

r.) $\cancel{2} \text{HgO} \longrightarrow \cancel{2} \text{Hg} + \text{O}_2$

s.) $\text{C}_2\text{H}_5\text{CO}_2 + \cancel{2} \text{O}_2 \longrightarrow \cancel{2} \text{CO}_2 + \cancel{2} \text{H}_2\text{O}$

t.) $\text{C}_2\text{H}_6 + \cancel{2} \text{O}_2 \longrightarrow \cancel{2} \text{CO}_2 + \cancel{2} \text{H}_2\text{O}$

u.) $\cancel{2} \text{FeCl}_3 + \cancel{2} \text{NH}_4\text{HS} \longrightarrow \cancel{2} \text{NH}_4\text{Cl} + \text{Fe}_2\text{S}_3$

v.) $\text{P}_2\text{O}_5 + \cancel{2} \text{NaOH} \longrightarrow \cancel{2} \text{Na}_3\text{PO}_4 + \cancel{2} \text{H}_2\text{O}$

w.) $\cancel{2} \text{BaCl}_2 + \text{Al}_2(\text{SO}_4)_3 \longrightarrow \cancel{2} \text{BaSO}_4 + \cancel{2} \text{AlCl}_3$

x.) $\cancel{2} \text{NH}_3 + \cancel{2} \text{O}_2 \longrightarrow \cancel{2} \text{H}_2\text{O} + \cancel{2} \text{NO}$

y.) $\cancel{2} \text{Na}_2\text{O}_2 + \cancel{2} \text{H}_2\text{O} \longrightarrow \cancel{2} \text{NaOH} + \cancel{2} \text{O}_2$

z.) $\text{Sn}(\text{CHO}_3)_2 \longrightarrow \text{SnCl}_4 + \cancel{2} \text{O}_2$

Two Problems to be Completed on the Other Side of this Worksheet:

*P. 173 #45, 47, 53, 54, 56, 57, 58, 63, 65, 66, 67, 71

*(Although the directions may say to write an "unbalanced" equation, write a balanced equation for each.)

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