

Chemical Bonding

1. OCTET RULE (LEWIS-KOSSEL RULE)

Noble gases have 8 electrons in their outermost shell (complete octet) and outermost configuration is ns^2p^6 . Every atom has a tendency to complete its octet by losing or gaining or by sharing electrons.

1.1 Exceptions of Octet Rule

(i) **Electron-Deficient Compounds** : Compounds having less than 8 electrons in the outermost shell of the central atom. e.g. BF_3 , BeCl_2 , AlCl_3 .

(ii) **Expansion of Octet** : The compounds having more than 8 electrons in the outermost shell of the central atom. e.g PCl_5 , SF_6 .

(iii) **Transition Metal Ions** : Transition metal ions have 9 to 18 electrons in their outermost shell. e.g. Cr^{+3} , Mn^{+2} , Cu^{+} .

(iv) H and Li atoms attain configuration like He by gaining and losing one electron, respectively. e.g. H^{+1} , Li^{+1} .

(v) **Odd Electron Bond** : Compounds having unpaired electron are also exceptions to the octet rule. The molecules having odd electrons are NO_2 , NO_3^- , ClO_4^- , O_2^{+1} etc.

Ex.1 Octet rule is not followed in:

$$[1] \text{ Ti}^{+3} \quad [2] \text{ Na}^{+} \quad [3] \text{ Ca}^{+2} \quad [4] \text{ K}^{+1} \quad \text{Ans. [1]}$$

Ques. Tl^{+3} has 9 electron in its outermost shell.

Ex.2 Which of the following compounds has the central atom with incomplete octet ?

[1] NH_3 [2] BCl_3 [3] CCl_4 [4] PCl_5 Ans. [2]

Sol. B has 6 electron in BCl_3 .

Ex.3 An element forms cation by losing electron. The configuration of the cation may be like that of :

[1] inert gas [2] pseudo inert gas [3] inert pair [4] all above Ans. [4]

Sol. Inert gas — 8 electron in outermost shell. Pseudo inert gas — 18 electron in outermost shell.

1.2 Inert Pair Configuration

The heavy elements of IIIA, IVA, and VA groups form cations having charge equal to group number and two less than group number, e.g. Ti^{+3} , Sn^{+2} , Pb^{+2} . In these elements due to more penetration power of s orbitals, the atom generally does not lose e^- present in ns orbital, so only np electrons are lost by the atom. This effect is called "inert pair effect".

2. IONIC BOND

(i) Atoms get established by ion formation in formation of an ionic bond. One atom forms a cation by losing electrons and other forms anion by accepting electrons.

(ii) The octets of the atoms are completed by transfer of electrons.

(iii) Cations and anions are bonded by electrostatic force of attraction in ionic bond.

(iv) Ionic bond is formed between metals (cations) and nonmetals (anions).

(v) During the formation of an ionic bond, cation can attain one of the following configurations.

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