

Reactions of Amines and Epoxides

Preparation of Amines

Amines are colorless, basic, and have a strong, fishy odor.



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Regioselectivity and Stereochemistry

Bond Types, Strength, Length, Polarity and Dipole Moment

Free-radical halogenation



There are two types of bonding: ionic bonding and covalent bonding.

• **Ionic bonds** are formed by the transfer of electrons from one element to another (ionization potential) and electron affinity (electron).



• **Covalent bonds** are formed by sharing of electrons between two atoms.



The Relative Electronegativity of Atoms



Bond and Bonding



Bonding energy



There are two types of bonding

Definition of ionic bond



Sigma and pi bonds

Covalent bonds are classified into sigma and pi bonds.

• **Sigma bonds** are formed by head-on overlap of orbitals.

• **pi bonds** are formed by side-on overlap of orbitals.

• **One single bond** is a sigma bond.

• **Two multiple bonds** consist of one sigma and two pi bonds.



pi bonds are weaker than sigma bonds.

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity values, covalent bonds can be polar and nonpolar.

• **Polar covalent bond** is formed when the difference in electronegativity is between 0.5 and 1.7.

• **Nonpolar covalent bond** is formed when the difference in electronegativity is less than 0.5.

• **Electronegativity** is the ability of an atom to attract electrons towards itself. The electronegativity values for various elements are given in the table below.

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Dipole Moment

The polarity of a bond is described by the dipole moment (μ).

It is a vector quantity.



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• The dipole moment is defined as a vector pointing from the less electronegative to the more electronegative atom.

• The vector dipole moment is a product of the charge (q) and the distance (r) between the charges.

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Molecular Dipole

To determine the molecular polarity, the vector sum of the dipole moments is calculated.

The polarity of the molecule is determined by the vector sum of the dipole moments of the bonds. A good example is carbon dioxide.



Carbon dioxide is a linear molecule and the dipole moment is zero.

Bond Length and Bond Strength

The bond length is the distance between the nuclei of the two atoms.

Bond	H-H	H-Cl	H-Br	H-I
Length	109 pm	127 pm	141 pm	154 pm
Strength	436 kJ/mol	431 kJ/mol	366 kJ/mol	298 kJ/mol

Double bonds are stronger than single bonds and are shorter.

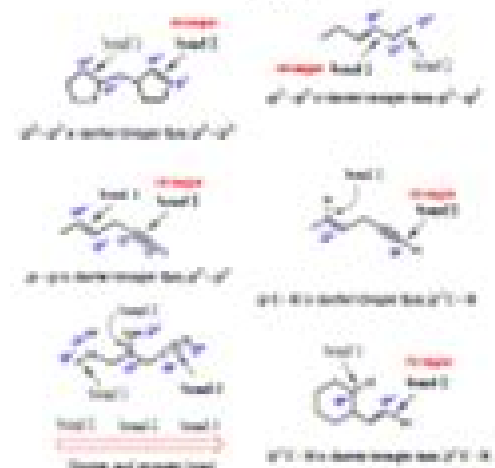


The bond length is the distance between the nuclei of the two atoms.

	Electronegativity	Electronegativity	Electronegativity
C, F	2.5	4.0	4.0
Electronegativity	2.5	4.0	4.0
Length	109 pm	127 pm	141 pm
Strength	436 kJ/mol	431 kJ/mol	366 kJ/mol
Electronegativity	2.5	4.0	4.0
Electronegativity	2.5	4.0	4.0
Length	109 pm	127 pm	141 pm
Strength	436 kJ/mol	431 kJ/mol	366 kJ/mol

Higher electronegativity values result in shorter bonds.

Examples



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