

Specific Heat Problems – ANSWERS

The last page of this document has Miscellaneous Information that you will need for these problems.

1. How many calories are needed to warm 55.8 g of water at 21.3°C to 73.4°C?

$$\text{Heat} = m \times S. H. \times (T_{\text{final}} - T_{\text{initial}})$$

$$\text{Heat} = 55.8 \text{ g} \times 1.00 \text{ cal/g}^\circ\text{C} \times (73.4^\circ\text{C} - 21.3^\circ\text{C})$$

$$\text{Heat} = 55.8 \text{ g} \times 1.00 \text{ cal/g}^\circ\text{C} \times 52.1^\circ\text{C}$$

$$\text{Heat} = 2907.18 \text{ cal} \quad \text{convert to scientific notation}$$

$$= 2.9\textbf{0}718 \times 10^3 \text{ cal}$$

$$= 2.91 \times 10^3 \text{ cal (3 sig. figs.)}$$

2. How many calories are needed to warm 55.8 g of alcohol from 21.3°C to 73.4°C?

$$\text{Heat} = m \times S. H. \times (T_{\text{final}} - T_{\text{initial}})$$

$$\text{Heat} = 55.8 \text{ g} \times 0.58 \text{ cal/g}^\circ\text{C} \times (73.4^\circ\text{C} - 21.3^\circ\text{C})$$

$$\text{Heat} = 55.8 \text{ g} \times 0.58 \text{ cal/g}^\circ\text{C} \times 52.1^\circ\text{C}$$

$$\text{Heat} = 1686.16 \text{ cal} \quad \text{convert to scientific notation}$$

$$\text{Heat} = 1.6\textbf{8}616 \times 10^3 \text{ cal}$$

$$\text{Heat} = 1.69 \times 10^3 \text{ cal}$$

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Table of Contents Organic Chemistry Specific Heat Problems

1. Understanding the eBook Organic Chemistry Specific Heat Problems
 - The Rise of Digital Reading Organic Chemistry Specific Heat Problems
 - Advantages of eBooks Over Traditional Books
2. Identifying Organic Chemistry Specific Heat Problems
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Organic Chemistry Specific Heat Problems
 - User-Friendly Interface
4. Exploring eBook Recommendations from Organic Chemistry Specific Heat Problems
 - Personalized Recommendations
 - Organic Chemistry Specific Heat Problems User Reviews and Ratings
 - Organic Chemistry Specific Heat Problems and Bestseller Lists
5. Accessing Organic Chemistry Specific Heat Problems Free and Paid eBooks
 - Organic Chemistry Specific Heat Problems Public Domain eBooks
 - Organic Chemistry Specific Heat Problems eBook Subscription Services
 - Organic Chemistry Specific Heat Problems Budget-Friendly Options
6. Navigating Organic Chemistry Specific Heat Problems eBook Formats

- ePub, PDF, MOBI, and More
- Organic Chemistry Specific Heat Problems Compatibility with Devices
- Organic Chemistry Specific Heat Problems Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Organic Chemistry Specific Heat Problems
 - Highlighting and Note-Taking Organic Chemistry Specific Heat Problems
 - Interactive Elements Organic Chemistry Specific Heat Problems
- 8. Staying Engaged with Organic Chemistry Specific Heat Problems
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Organic Chemistry Specific Heat Problems
- 9. Balancing eBooks and Physical Books Organic Chemistry Specific Heat Problems
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Organic Chemistry Specific Heat Problems
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Organic Chemistry Specific Heat Problems
 - Setting Reading Goals Organic Chemistry Specific Heat Problems
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Organic Chemistry Specific Heat Problems
 - Fact-Checking eBook Content of Organic Chemistry Specific Heat Problems
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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