

Reactions of Amines and Epoxides

Preparation of Amines

Amines are colorless, basic, and have a strong, fishy odor.



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Regioselectivity and Stereochemistry

Bond Types, Strength, Length, Polarity and Dipole Moment

Free-radical halogenation



There are two types of bonding: ionic bonding and covalent bonding.

• **Ionic bonds** are formed by the transfer of electrons from one element to another (ionization potential) and electron affinity (electron). Examples:



• **Covalent bonds** are formed by sharing of electrons between two atoms.



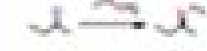
The Relative Electronegativity of Atoms



Bond and Bonding



Bonding energy



There are two types of electronegativity

Electronegativity of atoms is



Sigma and pi bonds

Covalent bonds are classified as sigma and pi bonds.

• **Sigma bonds** are formed by head-on (end-to-end) overlap of orbitals.

• **pi bonds** are formed by side-on overlap of orbitals.

• **One single bond is a sigma bond.**

• **Two multiple bonds consist of one sigma and two pi bonds.**



pi bonds are stronger than sigma bonds.

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity

values, covalent bonds can be polar and nonpolar.

• **Polar bonds** are formed by the difference in electronegativity between two atoms.

• **Nonpolar bonds** are formed by the difference in electronegativity between two atoms, forming a nonpolar covalent bond. The important examples are the C-C and C-H bonds.

• **If the difference in electronegativity is between 0.5 and 1.7, we have a polar covalent bond. The important examples are bonds for carbon to fluorine, N, O, F and other halogens.**

• **A difference of 1.7 or higher can be high that the electrons are not longer shared and a ionic bond is formed.**

Dipole Moment

The polarity of a bond is described by the dipole moment (μ).

It is a vector quantity.



The dipole moment is a vector pointing from the less electronegative to the more electronegative atom.

• The dipole moment is a vector pointing from the less electronegative to the more electronegative atom.

• The vector electronegativity of an atom is a vector pointing from the less electronegative to the more electronegative atom.

• **Electronegativity** indicates an electron-withdrawing atom.

• **Electronegativity** indicates an electron-donating atom.

Molecular Dipole

To determine the molecular polarity, the vector sum of the dipole moment.

The polarity of the molecule is a vector sum of the dipole moments of the atoms. The polarity of the molecule is a vector sum of the dipole moments of the atoms.



Overall the molecule is a polar molecule.



Overall the molecule is a nonpolar molecule.

Bond Length and Bond Strength

The bond length is the distance between the nuclei of the two atoms.

Bond	H-H	H-Cl	H-Br	H-I
Length	109 pm	127 pm	141 pm	154 pm
Strength	436 kJ/mol	431 kJ/mol	366 kJ/mol	298 kJ/mol

The bond length is the distance between the nuclei of the two atoms.



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	Electronegativity	Electronegativity	Electronegativity
C-F bond	4.0	3.5	2.5
C-Cl bond	3.0	2.5	1.5
C-Br bond	2.8	2.5	1.5
C-I bond	2.5	2.5	1.5
C-H bond	2.5	2.5	1.5
C-O bond	3.5	3.0	2.0
C-N bond	3.0	2.5	1.5
C-C bond	2.5	2.5	1.5
C-S bond	2.5	2.0	1.0
C-P bond	2.5	2.0	1.0
C-B bond	2.0	2.0	1.0
C-Mg bond	1.3	1.3	0.5
C-Na bond	0.9	0.9	0.5
C-K bond	0.8	0.8	0.5
C-Li bond	0.9	0.9	0.5
C-Al bond	1.5	1.5	0.5
C-Si bond	1.9	1.9	0.5
C-Ge bond	2.0	2.0	0.5
C-Sn bond	1.8	1.8	0.5
C-Pb bond	2.3	2.3	0.5
C-Bi bond	2.0	2.0	0.5
C-Tl bond	1.8	1.8	0.5
C-Po bond	2.0	2.0	0.5
C-At bond	2.2	2.2	0.5
C-Fl bond	4.0	3.5	2.5
C-Br bond	2.8	2.5	1.5
C-I bond	2.5	2.5	1.5
C-At bond	2.2	2.2	0.5
C-Po bond	2.0	2.0	0.5
C-Ge bond	2.0	2.0	0.5
C-Sn bond	1.8	1.8	0.5
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C-Tl bond	1.8	1.8	0.5
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C-At bond	2.2	2.2	0.5

The bond length is the distance between the nuclei of the two atoms.

Examples



Overall the molecule is a polar molecule.



Overall the molecule is a nonpolar molecule.



Overall the molecule is a polar molecule.



Overall the molecule is a nonpolar molecule.



Overall the molecule is a polar molecule.



Overall the molecule is a nonpolar molecule.

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Morrison



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