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Oracle Database 11g SQL

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Oracle Database 12c PL/SQL Programming

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web zero degrees celsius or 273 k is defined as standard temperature and one atmosphere or 760 mm hg is defined as

standard pressure the volume occupied by one mole of a gas is called molar volume therefore the molar volume of a gas at standard temperature and pressure stp is 22 4 l

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web jun 8 2023 $v = 1.00\text{mol} \times 0.08206 \text{latm} \text{molk}^{-1} \text{atm}^{-1} \text{mol}^{-1} \times 273\text{k} \times 1.00\text{atm} = 22.4\text{l}$ in other words 1 mole of a gas will occupy 22 4 l at stp

assuming ideal gas behavior at stp the volume of a gas is only dependent on number of moles of

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mole of any gas at room temperature and pressure occupies 24dm 3 24 000cm 3 this value is known as the molar volume

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