

Question 1

A 5.00 L sample of $\text{N}_2(\text{g})$ is held at a pressure of 2.00 atm and a temperature of 315 K. Assuming ideal-gas behavior, how many moles of $\text{N}_2(\text{g})$ are present in this sample?

- a. 2.69 mol
- b. 0.208 mol
- c. 0.387 mol
- d. 0.004 mol

Question 2

At a temperature of 500°C and a pressure of 699 Torr, sulfur vapor has a density of 3.71 g/dm³. What molecular formula for sulfur is compatible with this set of conditions?

- a. S_2
- b. S_3
- c. S_4
- d. (S_{25})

Question 3

Consider a gas for which the equation of state $P(V_m - b) = RT$ (with $b = 0.04 \text{ dm}^3/\text{mol}$) can be used to describe its P-V-T behavior. Suppose a 2-mol sample of this gas is compressed isothermally and reversibly from an initial volume of 10 dm³ to a final volume of 1 dm³. How much work (w) is done on the gas sample in this compression process, and how much is the internal energy (U) of the gas changed?

- a. $w = 11.7 \text{ kJ}$ and $\Delta U = 11.7 \text{ kJ}$
- b. $w = 11.9 \text{ kJ}$ and $\Delta U = 0$
- c. $w = 5.85 \text{ kJ}$ and $\Delta U = -5.85 \text{ kJ}$
- d. $w = 5.85 \text{ kJ}$ and $\Delta U = 0$

Question 4

Consider a process in which a sample of water vapor expands reversibly and adiabatically from a state in which its pressure is 60 Torr and its volume is 0.5 dm³ to a state in which its volume is 3 dm³. If it is assumed that the water vapor behaves as an ideal gas (with respect to its P-V-T properties) and that the ratio of its $C_{p,m}$ and $C_{v,m}$ heat-capacity quantities is given by $\gamma = C_{p,m}/C_{v,m} = 1.3$, what will be the final pressure of the vapor sample?

- a. 5.83 bar
- b. 230 Torr
- c. 5.84 Torr
- d. There is not enough information provided to answer this question.

Question 5

Consider a process in which exactly 107 J of heat is added to 1 mol of $\text{CH}_3\text{OH}(\text{g})$ under constant-volume conditions. If the initial temperature of the methanol vapor sample is 298 K, what will be the temperature of the vapor after the 107 J of heat is added? You may assume ideal-gas behavior and use a value of 43.89 J/K-mol for $C_{p,m}$ (the molar heat capacity of methanol vapor under constant-pressure conditions).

- a. 301.0 K
- b. 300.4 K
- c. 295.0 K
- d. None of the above answers is correct.

Question 6

If a person breathes at a rate of 14 respirations per minute, and each respiration involves 0.5 dm³ of air, what will be the total mass of air breathed in a day? (Take $T = 300 \text{ K}$ and $P = 101.3 \text{ kPa}$, and assume that the air behaves as an ideal gas with a molar mass of 0.029 kg/mol).

- a. 11.9 kg
- b. 119 g
- c. 238 g
- d. 14.6 kg

Question 7

If you have 1.55 mol of $\text{F}_2(\text{g})$ at 1.11 atm pressure and a temperature of 0°C, what is the volume of this gas sample, assuming ideal-gas behavior?

- a. 31.3 L
- b. 22.4 L
- c. 3170 L
- d. 0.0821 L

Question 8

The density of ice, $\text{H}_2\text{O}(\text{s})$, at 273 K, is 0.915 kg/dm³, and that of liquid water is 0.99987 kg/dm³. How much thermodynamic work must be done to melt 1 mol of ice at 1 bar pressure?

- a. 0.167 J
- b. 0.167 kJ
- c. 1.67 J
- d. 39.9 cal

Question 9

The heat capacity of a substance can be measured under either constant-volume or constant-pressure conditions. Which of the following sets of relationships between C_v (heat capacity under constant-volume conditions) and C_p (heat capacity under constant-pressure conditions) are always true?

- a. $C_p(\text{gas}) > C_v(\text{gas})$; $C_p(\text{solid}) = C_v(\text{solid})$; and $C_p(\text{gas}) >> C_p(\text{solid})$
- b. $C_v(\text{gas}) = C_v(\text{solid})$ and $C_p(\text{liquid}) = C_v(\text{liquid})$
- c. $C_v(\text{gas}) > C_v(\text{liquid}) > C_v(\text{solid})$ and $C_p > C_v$ in all phases
- d. C_p approaches a value of 0 as the sample temperature approaches 0 K but increases without bound as the sample temperature is increased.

Multiple Choice Physical Chemistry Exam Questions

John Moore, Richard H. Langley



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topics broadly can be grouped into Organic Inorganic and Physical chemistry Organic Chemistry is a vast section and holds the maximum weightage Its concepts and principles are important One needs to be conversant with properties of groups conversions recognizing organic compounds etc Interlinking between different chapters from organic chemistry can be of great help The students must primarily focus on reading NCERT textbooks Several times the questions asked in NEET exam are taken directly from the NCERT textbooks Initially avoid answering those questions for which you are not confident because your wrong answer may reduce your final score In order to utilize your time appropriately divide the three hours of examination time as per your comfort among Physics Chemistry and Biology Initially focus on attempting all easy questions and later on pick the difficult ones By this way your confidence will be elevated and you will also get more time to answer hard questions Practice previous years question papers mock tests and sample papers to get an idea on how to answer MCQ questions efficiently Preparing at an early stage is what an MCQ exam requires Avoid guesswork for negative marking questions as they might lower your final score These tips can be very helpful for students to answer difficult and brain teaser questions Prior preparations and practice are mandatory aspects of any examination Hence to crack highly competitive examination like NEET it is mandatory for students to prepare well and acquire the skills to tackle multiple choice questions effectively Rather than just following mere guesswork the aspirants can focus on the tips discussed to tackle Multiple Choice Questions in NEET in the right manner

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