

CHAPTER 6**STUDY GUIDE****Ionic Compounds and Metals****Lesson 6.1 Ion Formation**

In your textbook, read about chemical bonds and formation of ions.

Use each of the terms below just once to complete the passage.

| | | | | |
|---------------|-----------|-----------------------------|---------|-------------|
| chemical bond | electrons | energy level | ions | noble gases |
| nucleus | octet | pseudo-noble gas formations | valence | |

The force that holds two atoms together is called a(n) (1) **Chemical Bond**.

Such an attachment may form by the attraction of the positively charged

(2) **Nucleus** of one atom for the negatively charged

(3) **Electron** of another atom, or by the attraction of charged atoms, which are called (4) **Ions**. The attractions may also involve

(5) **Valence** electrons, which are the electrons in the outermost

(6) **Energy level**. The (7) **Noble Gas** are a family of elements that have very little tendency to react. Most of these elements have a set of eight outermost electrons, which is called a stable (8) **Octet**. The relatively stable electron structures developed by loss of electrons in certain elements of groups 3, 4, 13, and 14 are called (9) **Pseudo-Noble gas formations**.

For each statement below, write *true* or *false*.

False 10. A positively charged ion is called an anion.

True 11. Elements in group 1 lose their one valence electron, forming an ion with a $1+$ charge.

False 12. Elements tend to react so that they acquire the electron structure of a halogen.

True 13. A sodium atom tends to lose one electron when it reacts.

True 14. The electron structure of a zinc ion (Zn^{2+}) is an example of a pseudo-noble gas formation.

False 15. A Cl^- ion is an example of a cation.

True 16. The ending *-ide* is used to designate an anion.

False 17. Nonmetals form a stable outer electron configuration by losing electrons and becoming anions.

Metals Study Guide Answers

Brian Coffin, Kimberley Keller

Metals Study Guide Answers:

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