

Unit 1 Study Guide — ANSWER KEY

1. Define scientific method: **a procedure for collecting, testing, and communicating information**
2. Define hypothesis: **a suggested answer to the question, an educated guess**
3. Define experiment: **a scientific test so see what will happen**
4. Define dependent variable: **the variable that you do NOT control. (on a graph is usually on the Y axis)**
5. Define independent variable: **the variable that I control. (on a graph is usually on the X axis)**
6. Define control: **the part of the experiment where everything is left the same.**
7. Define observation: **information collected with any of your 5 senses**
8. Define density: **the ratio of mass per unit volume (D = m/v)**
9. Define accuracy: **the measure of how close your data is to the ACTUAL or true value.**
10. Define precision: **a measure of how close your data (measurements) are to each other.**
11. Why do scientists use SI units of measurement? **So that measurements can be understood internationally (around the world).**
12. List (in the correct order) the steps of the scientific method.

- | | |
|-------------------------------|--|
| 1. make an observation | 4. test hypothesis (experiment) |
| 2. ask a question | 5. analyze data/form a conclusion |
| 3. form a hypothesis | 6. develop a theory |

13. What mnemonic device is used to remember the metric system?

K H Da B d c m
g, m, L

14. What is the sentence (mnemsonic) that will help you remember the METRIC prefixes in order?

King Henry Died By Drinking Chocolate Milk

15. Fill in the table below describing the metric system:

	Kilo	Hecto	Deca	Base	Deci	Centi	Milli
Abbreviation	K	H	Da	g, m, L	d	c	m
Meaning in powers of 10	10^3	10^2	10^1	10^0	10^{-1}	10^{-2}	10^{-3}

Convert the following:

- | | |
|---------------------------------|------------------------------------|
| 16. 25L to <u>25000</u> mL | 21. 25 cm to <u>0.25</u> m |
| 17. 2.3 km to <u>230,000</u> cm | 22. 56784 mm to <u>0.056784</u> km |
| 18. 0.045 g to <u>45</u> mg | 23. 15 g to <u>0.015</u> kg |
| 19. 0.35 kg to <u>350</u> g | 24. 3.452 kL to <u>3452</u> L |
| 20. 300 mL to <u>0.300</u> L | 25. 47 m to <u>47000</u> mm |

Scientific Notation:

Convert the following to scientific notation:

- | | | | |
|------------|--------------------------------------|------------|--|
| 26) 0.005 | <u>5×10^{-3}</u> | 30) 0.025 | <u>2.5×10^{-2}</u> |
| 27) 5.050 | <u>5.05×10^1</u> | 31) 0.0025 | <u>2.5×10^{-3}</u> |
| 28) 0.0008 | <u>8×10^{-4}</u> | 32) 500 | <u>5×10^2</u> |
| 29) 1,000 | <u>1×10^3</u> | 33) 5,000 | <u>5×10^3</u> |

Pbs Unit 5 Study Guide Answer Key

R. Keith Sawyer



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web balancing chemical equations answer key balance the equations below $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$ $\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$ $2\text{NaCl} + \text{F}_2 \rightarrow 2\text{NaF} + \text{Cl}_2$ $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$ $\text{Pb} + \text{OH}^- \rightarrow 2\text{HCl} + 2\text{H}_2\text{O}$ $1\text{PbCl}_2 + 3\text{AlBr}_3 + 3\text{K}_2\text{SO}_4 \rightarrow 6\text{KBr} + 1\text{Al}_2\text{SO}_4 + 3\text{CH}_4 + 2\text{O}_2 \rightarrow 1\text{CO}_2 + 2\text{H}_2\text{O}$

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web may 1 2013 balancing chemical equations a balanced equation is a chemical equation in which mass is conserved and there are equal numbers of atoms of each element on both sides of the equation we can write a chemical equation for the

reaction of carbon with hydrogen gas to form methane CH_4

20 balancing chemical equations worksheets answers - Nov 30 2022

web a balancing chemical equation worksheet is a practice booklet with unsolved and solved chemical equation problems on which students can practice their balancing skills if all the problems are unsolved it will usually contain an answer section that students can use to gauge their performance

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$1 \text{ N}_2 + 3 \text{ H}_2 \rightarrow 2 \text{ NH}_3$ $3 \text{ K}_2\text{O} + 2 \text{ KClO}_3 \rightarrow 6 \text{ KCl} + 3 \text{ O}_2$ $2 \text{ NaCl} + 1 \text{ F}_2 \rightarrow 2 \text{ NaF}$ $1 \text{ Cl}_2 + 4 \text{ H}_2 \rightarrow 2 \text{ H}_2\text{O} + 2 \text{ H}_2$ $5 \text{ Pb} + 2 \text{ OH}^- + 2 \text{ HCl} \rightarrow 2 \text{ H}_2\text{O} + 1 \text{ PbCl}_2 + 6 \text{ Al} + 3 \text{ K}_2\text{SO}_4 + 6 \text{ KBr} + 1 \text{ Al}_2\text{SO}_4 + 7 \text{ H}_2\text{O} + 1 \text{ CH}_4 + 2 \text{ O}_2 + 1 \text{ CO}_2 + 2 \text{ H}_2\text{O} + 8 \text{ C} + 3 \text{ H}_2 + 8 \text{ O}_2 + 3 \text{ CO}_2 + 4$

5 1 1 practice problems writing and balancing chemical equations - May 05 2023

web may 28 2020 write a balanced chemical equation for each step of the process the first step is the decomposition of solid calcium carbonate from seashells to form solid calcium oxide and gaseous carbon dioxide the second step is the formation of solid calcium hydroxide as the only product from the reaction of the solid calcium oxide with liquid water