

Reactions of Amines and Epoxydes

Preparation of Amines

Ammonium hydroxide: $\text{NH}_4^+ + \text{OH}^- \rightarrow \text{NH}_3 + \text{H}_2\text{O}$



Four different types of bonds



There are four types of bonding: ionic bonding and covalent bonding.

- Ionic bonds are formed by the transfer of electrons from one atom to another (anions and cations) (opposite charges).



Covalent ionic and polar covalent bonds

- Covalent bonds are classified as ionic (polar) and non-polar.
- Ionic bonds are formed by the transfer of electrons.
- Non-polar bonds are formed by the sharing of electrons.
- Polar covalent bonds are formed by the sharing of electrons.
- Non-polar covalent bonds are formed by the sharing of electrons.



Non-polar covalent bonds are stronger than ionic.

Strength of different bonds

Ionic bonds



Covalent bonds



Polar covalent bonds



Non-polar covalent bonds



Polar and Nonpolar Covalent Bonds



- If the difference in electronegativity is less than 0.7, the electrons are shared equally between the two atoms. Forming a nonpolar covalent bond. (Nonpolar covalent bonds are the H-H and Cl-Cl bonds.)
- If the difference in electronegativity is between 0.3 and 1.7, the more polar covalent bond. (More covalent bonds in one atom than the other in H-F, H-Cl, H-Br, and H-I bonds.)
- A difference of 1.7 or higher one or both of the atoms will have a larger share and a polar bond is formed.

A Summary of Ring-Opening Reactions of Epoxydes with Strong and Weak Protagonists

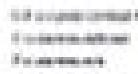


Reactions and Mechanisms

Bond Types, Strength, Length, Polarity and Dipole Moment

Dipole Moment

The polarity of a bond is denoted by the dipole moment (μ).



Directional dipole moment represents the polarity of the bond.

- The dipole moment is shown as a vector pointing from the less electronegative atom to the more electronegative atom.
- The vector represents the size for a partial negative charge and the less electronegative atom a partial positive charge.
- The dipole plot indicates the charge difference.
- The dipole moment indicates the charge and size.

Molecular Dipole

To measure the molecular polarity, the sum of the dipole is used.

The polarity of the molecule is equal to the sum of the bond dipole moments. The dipole moment is measured. A good example is water, dipole.



Sum of vector of the bond dipole moment for the molecule.

Bond Length and Bond Strength

Reactions of bond length



Bond length = 1.77 \AA (H-Cl bond)

Bond length = 1.47 \AA (H-Br bond)

Bond length = 1.37 \AA (H-I bond)

Bond length = 1.27 \AA (H-F bond)

Bond length = 1.16 \AA (H-O bond)

Bond length = 1.01 \AA (H-N bond)

Bond length = 0.74 \AA (H-C bond)

Bond length = 0.66 \AA (H-H bond)

Reactions of bond strength

Average values of C-C and C-H bond depending on hybridization

C-C Bond	$\text{C}_2\text{H}_2-\text{CH}_3$	$\text{C}_2\text{H}_5-\text{CH}_3$	$\text{C}_6\text{H}_5-\text{CH}_3$
Hybridization	sp^2-sp^2	sp^3-sp^2	sp^2-sp
Length	1.77 \AA	1.47 \AA	1.37 \AA
Strength	100 kJ/mol	112 kJ/mol	100 kJ/mol
Hybridization	sp^2	sp^3	sp^2
Efficiency	25%	33%	33%
Length	1.16 \AA	1.01 \AA	0.74 \AA
Strength	100 kJ/mol	112 kJ/mol	100 kJ/mol

Higher the bond length and stronger bond

Example



$\text{H}-\text{Cl} \text{ bond length} = 1.77 \text{ \AA}$

$\text{H}-\text{Br} \text{ bond length} = 1.47 \text{ \AA}$

$\text{H}-\text{I} \text{ bond length} = 1.37 \text{ \AA}$

$\text{H}-\text{F} \text{ bond length} = 1.27 \text{ \AA}$

$\text{H}-\text{O} \text{ bond length} = 1.16 \text{ \AA}$

$\text{H}-\text{N} \text{ bond length} = 1.01 \text{ \AA}$

$\text{H}-\text{C} \text{ bond length} = 0.74 \text{ \AA}$

$\text{H}-\text{H} \text{ bond length} = 0.66 \text{ \AA}$

Properties of covalent bonds

Pavia Organic Chemistry Lab Study Guide

Emilie Sanchez

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