

Nomenclature of Ionic Compounds (ANSWER KEY)

Binary Ionic Compounds:

Writing Names	Writing Formula
1. Write the name of the cation (metal) 2. Write the name of the anion (non-metal), drop the last syllable of the name, and change the suffix to "-ide". (Note: for the anions of O, N, H, P, Se, and Te drop the last two syllables)	1. Examine the name of the compound. Identify the ions and their oxidation states. 2. Determine the number of each ion needed to balance the charges. 3. Note the ratio of positive to negative ions, and write the formula. 4. Write the chemical formula, using subscripts if needed. (Note: for ionic compounds reduce the subscripts whenever possible.)

Complete the following table for binary ionic compounds.

Chemical Name	Chemical Formula	Chemical Name	Chemical Formula
sodium fluoride	NaF	beryllium fluoride	BeF ₂
lithium chloride	LiCl	sodium iodide	NaI
beryllium bromide	BeBr ₂	magnesium oxide	MgO
magnesium oxide	MgO	**Omitted**	H ₂ O
boron iodide	BI ₃	lithium sulfide	Li ₂ S
aluminum sulfide	Al ₂ S ₃	**Omitted**	BCl ₃
potassium oxide	K ₂ O	aluminum oxide	Al ₂ O ₃
calcium fluoride	CaF ₂	potassium sulfide	K ₂ S
barium nitride	Ba ₃ N ₂	silver chloride	AgCl
cesium sulfide	Cs ₂ S	calcium chloride	CaCl ₂
strontium oxide	SrO	**Omitted**	BN
francium bromide	FrBr	cesium iodide	CsI

Binary Ionic Compounds with Multivalent Cations:

Recall that multivalent elements have multiple oxidation states (charges). When writing the names for compounds containing multivalent cations we must indicate the oxidation state of the cation.

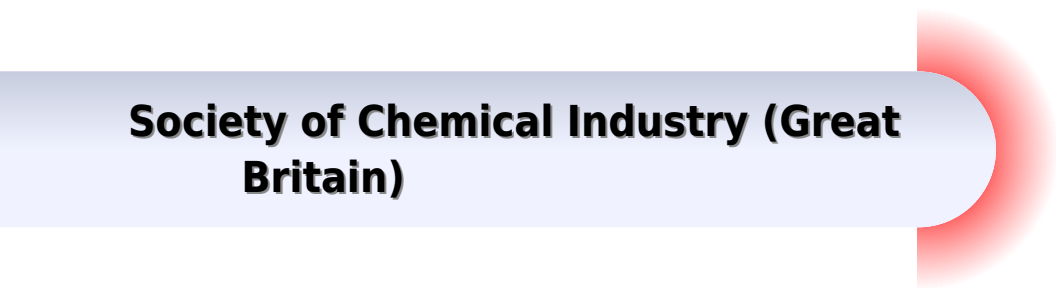
There are two methods for naming binary ionic compounds with multivalent cations.

1. **The classical system** (also referred to as the ous-ic method)
2. **The IUPAC system** (also referred to as the roman numeral method)

Naming Molecular Compounds Note Taking Guide

Answer

Society of Chemical Industry (Great Britain)



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