

AP Chemistry Pacing Guide

Unit 4: Chemical Reactions <ul style="list-style-type: none"> - Introduction for Reactions - Net Ionic Equations - Representations of Reactions - Physical and Chemical Changes - Stoichiometry - Introduction to Titration - Types of Chemical Reactions - Introduction to Acid-Base Reactions - Oxidation-Reduction (Redox) Reactions 	BIG IDEA 1c: Scale, Proportion, and Quantity BIG IDEA 3c: Transformations	<ol style="list-style-type: none"> I can identify evidence of chemical and physical changes in matter. I can represent changes in matter with a balanced chemical or net ionic equation: a. For physical changes. b. For given information about the identity of the reactants and/or product. For ionic in a given chemical reaction I can represent a given chemical reaction or physical process with a consistent particulate model. I can explain the relationship between macroscopic characteristics and bond interactions for: a. Chemical processes. b. Physical processes. I can explain changes in the amounts of reactants and products based on the balanced reaction equation for a chemical process. I can identify the equivalence point in a titration based on the amounts of the titrant and analyte, assuming the titration reaction goes to completion. I can identify a reaction as acid-base, oxidation-reduction, or precipitation. I can identify species as Brønsted-Lowry acids, bases, and/or conjugate acid-base pairs. I can represent a balanced redox reaction equation using half-reactions. 	10 day(s)
Unit 5: Kinetics <ul style="list-style-type: none"> - Reaction Rates - Introduction to Rate Law - Concentration Changes Over Time - Elementary Reactions - Collision Model - Reaction Energy Profile - Introduction to Reaction Mechanisms - Mechanisms Reaction Mechanism and Rate Law - Steady State 	BIG IDEA 3c: Transformations BIG IDEA 4i: Energy	<ol style="list-style-type: none"> I can explain the relationship between the rate of a chemical reaction and experimental parameters. 2. I can represent experimental data with a consistent rate law expression. I can identify the rate law expression of a chemical reaction using data that show how the concentrations of reaction species change over time. I can represent an elementary reaction as a rate law expression using stoichiometry. I can explain the relationship between the rate of an elementary reaction and the frequency, energy, and orientation of molecular collisions. I can represent the activation energy and overall energy change in an elementary reaction using a reaction energy profile. I can identify the components of a reaction mechanism. I can identify the rate law for a reaction from a mechanism in which the first step is rate limiting. I can identify the rate law for a reaction from a mechanism in which the 	10 day(s)

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