

# AP Chemistry Pacing Guide

<b>Unit 4: Chemical Reactions</b> <ul style="list-style-type: none"> <li>Introduction for Reactions</li> <li>Net Ionic Equations</li> <li>Representations of Reactions</li> <li>Physical and Chemical Changes</li> <li>Stoichiometry</li> <li>Introduction to Titration</li> <li>Types of Chemical Reactions</li> <li>Introduction to Acid-Base Reactions</li> <li>Oxidation-Reduction (Redox) Reactions</li> </ul>	<b>BIG IDEA 1:</b> Scale, Proportion, and Quantity <b>BIG IDEA 3:</b> Transformations	<ol style="list-style-type: none"> <li>I can identify evidence of chemical and physical changes in matter.</li> <li>I can represent changes in matter with a balanced chemical or net ionic equation: a. For physical changes; b. For given information about the identity of the reactants and/or product.</li> <li>For ions in a given chemical reaction I can represent a given chemical reaction or physical process with a consistent particulate model.</li> <li>I can explain the relationship between macroscopic characteristics and bond interactions for: a. Chemical processes; b. Physical processes.</li> <li>I can explain changes in the amounts of reactants and products based on the balanced reaction equation for a chemical process.</li> <li>I can identify the equivalence point in a titration based on the amounts of the titrant and analyte, assuming the titration reaction goes to completion.</li> <li>I can identify a reaction as acid-base, oxidation-reduction, or precipitation.</li> <li>I can identify species as Brønsted-Lowry acids, bases, and/or conjugate acid-base pairs, based on proton-transfer involving those species.</li> <li>I can represent a balanced redox reaction equation using half-reactions.</li> </ol>	10 days
<b>Unit 5: Kinetics</b> <ul style="list-style-type: none"> <li>Reaction Rates</li> <li>Introduction to Rate Law</li> <li>Concentration Changes Over Time</li> <li>Elementary Reactions</li> <li>Collision Model</li> <li>Reaction Energy Profile</li> <li>Introduction to Reaction Mechanism</li> <li>Mechanism Reaction Mechanism and Rate Law</li> <li>Steady-State</li> </ul>	<b>BIG IDEA 3:</b> Transformations <b>BIG IDEA 6:</b> Energy	<ol style="list-style-type: none"> <li>I can explain the relationship between the rate of a chemical reaction and experimental parameters.</li> <li>I can represent experimental data with a consistent rate law expression.</li> <li>I can identify the rate law expression of a chemical reaction using data that show how the concentrations of reaction species change over time.</li> <li>I can represent an elementary reaction as a rate law expression using stoichiometry.</li> <li>I can explain the relationship between the rate of an elementary reaction and the frequency, energy, and orientation of molecular collisions.</li> <li>I can represent the activation energy and overall energy change in an elementary reaction using a reaction energy profile.</li> <li>I can identify the components of a reaction mechanism.</li> <li>I can identify the rate law for a reaction from a mechanism in which the first step is rate limiting.</li> <li>I can identify the rate law for a reaction from a mechanism in which the</li> </ol>	10 days

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