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**Rachel S Tattersall**



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web there's 2 on the left and 1 on the right so we need to change the coefficient of  $\text{NH}_3$  to 2 now we have  $\text{N}_2 + \text{H}_2 \rightarrow 2 \text{ NH}_3$  total the atoms up again on the left there is 2 N and 2 H still on the right there is 2 N and 6 H now so now all we need to do is make the left side have 6 H in total

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oxygen gas to produce water in the gaseous phase was shown as a chemical equation  $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$  at the molecular level the reaction would look something like this notice that there are two oxygen atoms on the

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web in order to balance a chemical equation the quantities of each type of element and polyatomic ion that are present in the reactants and the products of the reaction must be determined the first reactant in this equation contains magnesium mg and a hydroxide ion oh<sup>-</sup> which is a polyatomic anion

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