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**Jayant Deshmukh**



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web if you start with 14.82 g of  $\text{Ca(OH)}_2$  and 16.35 g of  $\text{H}_2\text{SO}_4$  a determine the limiting reagent b determine the number of moles of  $\text{H}_2\text{O}$  produced c determine the number of grams of  $\text{CaSO}_4$  produced d determine the number of grams of excess reagent left 1 make sure the equation is balanced

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