

KEY

1. Complete and classify the following partial equations

- $4\text{Fe} + 6\text{H}_2\text{SO}_4 \rightarrow 2\text{Fe}_2\text{SO}_4 + 4\text{H}_2$ (single rep) $2\text{Fe} + 3\text{H}_2\text{SO}_4 \rightarrow \text{Fe}_2(\text{SO}_4)_3 + 3\text{H}_2$
- $\text{Na} + \text{Cl}_2 \rightarrow$
- $12\text{Zn} + 4\text{Pb}(\text{NO}_3)_2 \rightarrow 12\text{Zn}(\text{NO}_3)_2 + 4\text{Pb}$ (single rep)
- $\text{Al}_2\text{O}_3 + \text{HCl} \rightarrow$
- $2\text{Cu} + \text{O}_2 \rightarrow 2\text{CuO}$ $4\text{Cu} + \text{O}_2 \rightarrow 2\text{Cu}_2\text{O}$ (synthesis)
- $\text{Al}_2\text{O}_3 \rightarrow$
- $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$ (synthesis)
- $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow$
- $1\text{Mg} + 2\text{HNO}_3 \rightarrow \text{Mg}(\text{NO}_3)_2 + \text{H}_2$ (c.r.)
- $\text{Na} + \text{H}_2\text{O} \rightarrow$
- $3\text{NaOH} + \text{H}_3\text{PO}_4 \rightarrow \text{Na}_3\text{PO}_4 + 3\text{H}_2\text{O}$ (acid base)
- $\text{H}_2\text{SO}_4 + \text{Al}(\text{OH})_3 \rightarrow$
- $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$ (combustion)
- $\text{AlCl}_3 \rightarrow$
- $2\text{KOH} + \text{H}_2\text{SO}_4 \rightarrow \text{K}_2\text{SO}_4 + 2\text{H}_2\text{O}$ (acid base)
- $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow$
- $\text{NH}_4\text{OH} + \text{HCl} \rightarrow \text{NH}_4\text{Cl} + \text{H}_2\text{O}$ (acid base)
- $\text{NaHCO}_3 + \text{H}_2\text{SO}_4 \rightarrow$
- $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$ (combustion)
- $\text{H}_2\text{O} \rightarrow$
- $2\text{HNO}_3 + \text{CaCO}_3 \rightarrow \text{Ca}(\text{NO}_3)_2 + \text{H}_2\text{O} + \text{CO}_2$ (c.r.)
- $\text{CH}_4 + \text{O}_2 \rightarrow$
- $2\text{Pb} + 4\text{H}_2\text{SO}_4 \rightarrow 2\text{PbSO}_4 + 4\text{H}_2$ (c.r.)
- $\text{KI} + \text{Br}_2 \rightarrow$
- $\text{Ba}(\text{OH})_2 + 2\text{HNO}_3 \rightarrow \text{Ba}(\text{NO}_3)_2 + 2\text{H}_2\text{O}$ (acid base)
- $\text{Pb} + \text{Cu}(\text{NO}_3)_2 \rightarrow$

2. Given: $2\text{NaClO}_3 \rightarrow 2\text{NaCl} + 3\text{O}_2$ calculate the mass of oxygen gas produced if 143 g of NaClO_3 is decomposed.

65.4g

3. Calculate the mass of oxygen gas required to burn 242 g of wax ($\text{C}_{21}\text{H}_{44}$)

254g

4. If 14.0 moles of butane (C_4H_{10}) is burned

- calculate the number of moles of oxygen gas required. 98.0 mol O_2
- calculate the number of moles of carbon dioxide produced. 56.0 mol CO_2
- calculate the mass of water produced. 1080 mol H_2O

5. Metallic aluminum reacts with zinc nitrate to produce metallic zinc and aluminum nitrate.

- Calculate the mass of aluminum that will react with 1.30 moles of zinc nitrate. 27.4g Al
- Calculate the mass of zinc that would be produced from 50.0 g of aluminum and sufficient zinc nitrate. 182g Zn

6. Given: $\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu}(\text{NO}_3)_2 + 2\text{Ag}$

- How many grams of Cu are required to react with 3.00 g of silver nitrate? 0.544g Cu
- How many grams of silver are produced from 3.00 g of silver nitrate and sufficient Cu? 1.70g Ag

2. NaClO_3 143g
 NaClO_3 143g
 $143\text{g} \times \frac{1}{106.44} = 1.344\text{ mol}$
 $1.344 \text{ mol} \times \frac{3}{2} = 2.016 \text{ mol } \text{O}_2$
 $2.016 \text{ mol} \times 32 = 64.512\text{g}$
 65.4g

3. $\text{C}_{21}\text{H}_{44} + 32\text{O}_2 \rightarrow 21\text{CO}_2 + 22\text{H}_2\text{O}$
 $242\text{g} \times \frac{1}{254.48} = 0.951\text{ mol}$
 $0.951 \text{ mol} \times 32 = 30.432\text{g}$

Nelson Chemistry 30 Answer Key

Arthur James Wells



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