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language education This 15 chapter anthology examines the pedagogical ethical and cognitive aspects of AI enhanced instruction utilizing research from Vietnam and global contexts Section I Chapters 2 8 AI Integration in Language Teaching and Learning analyzes the role of AI in language instruction and acquisition encompassing emotional intelligence rubric based assessment IELTS preparation and educator identification Section II Chapters 9 13 AI Enhanced Language Learning Skills and Practices examines AI augmented language learning methodologies addressing vocabulary acquisition learner independence and chatbot assisted education Section III Chapters 14 15 AI in Translation and Support Systems examines artificial intelligence in translation and academic assistance systems The book emphasizes the transformative potential as well as the dangers of excessive dependence academic misconduct and insufficient teacher training Highlighting ethical integration and human AI collaboration it offers pragmatic ideas for educators academics and policymakers Empowering educators envisions AI as a strategically augmenting not supplanting teachers in the establishment of sustainable and inclusive language learning settings

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web donor whereas a base generically represented by B is a proton acceptor when acids and bases combine they swap a proton $H^+ + B \rightleftharpoons HB$ strong acids e.g. sulfuric acid have a very pronounced tendency to transfer a proton to a base whereas weak acids

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web understand that parts of an acid base titration be able to determine the K_a or K_b from pH data associated with the titration of a weak acid or base be able to determine the molar mass of a solid monoprotic acid from titration data be able to calculate K_{a1} and K_{a2} for a polyprotic acid by the end of this lab students should be able to

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completely reacts with the acid and the solution in the erlenmeyer flask continues to be acidic but at the equivalence point
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web answer based on the data the strong acid strong base combination of hydrochloric acid and sodium hydroxide had the neutral ph at the equivalence point being 7 this result is expected because a combination of strong acids and strong bases should result in a neutral at the equivalence point

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web 1 approximately 3 g of hydrated oxalic acid $\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$ was weighted to the nearest 0.01 g in a 50 ml beaker 2 about 30 ml of distilled water was added to the beaker to dissolve the oxalic acid 3 the solution was transferred into 250 ml volumetric flask the beaker was rinsed and its contents were poured into the flask

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