

Reactions of Amines and Epoxydes

Preparation of Amines

Ammonium hydroxide: $\text{NH}_4^+ + \text{OH}^- \rightarrow \text{NH}_3 + \text{H}_2\text{O}$



Four different types of bonds



There are four types of bonding: ionic bonding and covalent bonding.

- Ionic bonds are formed by the transfer of electron from one atom to another (anions and cations) (opposite charges).



- Covalent bonds are formed by sharing of electrons between two atoms.



Single (σ) and pi (π) bonds

- Covalent bonds are classified as single (σ) and pi (π).
- Single bonds are formed by head-on overlap of orbitals.
- π bonds are formed by side-on overlap of orbitals.
- Very simple bond is a covalent bond.
- Non-bonding bond consists of one sigma and two pi bonds.



π bond is stronger than σ bond.

Definition of Amino acid

Properties of Amino acid



A Summary of Ring-Opening Reactions of Epoxydes with Strong and Weak Nucleophiles

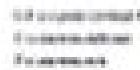


Reactions and Mechanisms

Bond Types, Strength, Length, Polarity and Dipole Moment

Dipole Moment

The polarity of a bond is denoted by the dipole moment (μ).



Directional dipole moment of the covalent bond.

- The dipole moment is shown as a vector pointing from the less electronegative bond to the more electronegative atom.
- The vector (dipole moment) starts from a partial negative charge and the vector (dipole moment) ends in a partial positive charge.
- The dipole plot indicates the charge difference.
- The dipole moment indicates the charge and area.

Molecular Dipole

To measure the molecular polarity, the sum of the dipole is used.

The polarity of the molecule is equal to the sum of the bond dipole moments. If the bond dipole moments are equal, it is just simple to calculate.



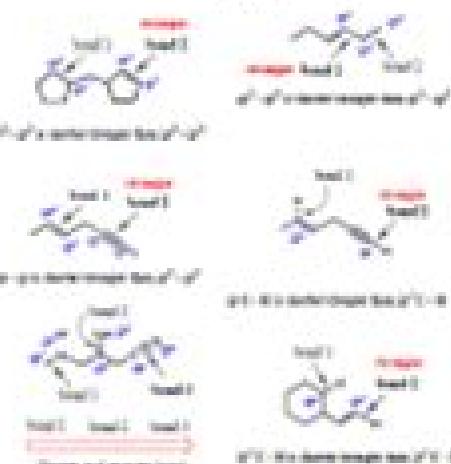
Sum of vector of the bond dipole for the water.

Average values of C-C and C-H bond depending on hybridization

C-C Bond	$\text{C}_2\text{H}_2-\text{CH}_3$	$\text{C}_2\text{H}_5-\text{CH}_3$	$\text{C}_6\text{H}_5-\text{CH}_3$
Hybridization	sp^2-sp^2	sp^3-sp^2	sp^2-sp
Length	1.37 Å	1.55 Å	1.39 Å
Strength	1.01 kcal/mol	0.77 kcal/mol	0.77 kcal/mol
Hybridization	sp^2	sp^3	sp^2
Efficiency	25.7%	33.7%	33.7%
Length	1.39 Å	1.53 Å	1.39 Å
Strength	0.91 kcal/mol	0.71 kcal/mol	0.71 kcal/mol

Dipole moment vector and strong bond

Example



Bond Length and Bond Strength

Relative bond length



Bond length: $\text{C}_2\text{H}_2-\text{CH}_3 < \text{C}_2\text{H}_5-\text{CH}_3 < \text{C}_6\text{H}_5-\text{CH}_3$

Strength: $\text{C}_2\text{H}_2-\text{CH}_3 > \text{C}_2\text{H}_5-\text{CH}_3 > \text{C}_6\text{H}_5-\text{CH}_3$

Efficiency: $\text{C}_2\text{H}_2-\text{CH}_3 > \text{C}_2\text{H}_5-\text{CH}_3 > \text{C}_6\text{H}_5-\text{CH}_3$

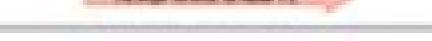
Length and strength: $\text{C}_2\text{H}_2-\text{CH}_3 > \text{C}_2\text{H}_5-\text{CH}_3 > \text{C}_6\text{H}_5-\text{CH}_3$

Relative bond strength



Bond strength: $\text{C}_2\text{H}_2-\text{CH}_3 > \text{C}_2\text{H}_5-\text{CH}_3 > \text{C}_6\text{H}_5-\text{CH}_3$

Relative bond to water



Bond to water: $\text{C}_2\text{H}_2-\text{CH}_3 > \text{C}_2\text{H}_5-\text{CH}_3 > \text{C}_6\text{H}_5-\text{CH}_3$

Properties of the molecule

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