



## Biology

### DNA Replication: Paper Clip Activity

Name \_\_\_\_\_ Per. \_\_\_\_\_ Date: \_\_\_\_\_

Partner: \_\_\_\_\_

Partner: \_\_\_\_\_

#### Learning Target: *DNA Structure & Replication*

***I Can...*** Create a simulated gene segment of DNA and take the segment through the process of replication.

***I Will...***

- Create a simulated primary segment of a gene representing the nucleotide types within the sequence with assigned colored paperclips.
- Predict and create a complementary strand of DNA using the base pairing rules.
- Unzip and replicate the DNA gene segment explaining the steps of the process.
- Infer and predict what may result from creating an alteration in the nucleotide sequence.

#### Quick Review:

- Each DNA molecule has a unique structure that makes it different from other DNA molecules (Remember – A chromosome is condensed DNA and segments of DNA are genes.)
- This difference occurs because the sequence of A, T, C, and G vary from one molecule and gene segment to another.



#### What You Need to Know About DNA Replication:

- To “replicate” DNA means to produce an exact copy of itself. DNA is the only known molecule that can do this regularly – RNA has been known to replicate under laboratory conditions.
- DNA is able to make an exact replica of itself because of the base pairing characteristics stressed earlier (A with T and C with G).
- When DNA makes a duplicate molecule of itself, the two strands unwind.
- After the two strands have pulled apart, new bases (A, T, C, & G) as well as new sugar and phosphate units come into place according to the base pairing rules.
- A comes in opposite of T, and C is opposite of G.
- When this occurs, two identical DNA molecules are created.



# Paper Clip Dna Replication Activity Answers

**Kathleen B. Sloan**



## **Paper Clip Dna Replication Activity Answers:**

*Cr 9 DNA* Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2004      *The Sourcebook for Teaching Science, Grades 6-12* Norman Herr, 2008-08-11 The Sourcebook for Teaching Science is a unique comprehensive resource designed to give middle and high school science teachers a wealth of information that will enhance any science curriculum Filled with innovative tools dynamic activities and practical lesson plans that are grounded in theory research and national standards the book offers both new and experienced science teachers powerful strategies and original ideas that will enhance the teaching of physics chemistry biology and the earth and space sciences      *Holt Biology*, 2004      *Mosby's Anatomy and Physiology Laboratory Manual* Kathleen B. Sloan, 1990      Biological Science Biological Sciences Curriculum Study, 1996      Thinking Inside the Block Schedule Pamela Robbins, Gayle H. Gregory, Lynne E. Herndon, 2000-04-20 This book is full of practical instructional strategies to help foster high levels of student achievement in the block schedule It contains strategies for differentiation powerful brain based teaching techniques creative approaches to productively organizing extended periods of time and proactive classroom management tips It adds to the repertoire every teacher needs to assure no child is left behind in the teaching learning process      *Mandell, Douglas, and Bennett's Principles and Practice of Infectious Diseases* John E. Bennett, MD, MACP, Raphael Dolin, MD, Martin J. Blaser, MD, 2014-08-28 After thirty five years Mandell Douglas and Bennett s Principles and Practice of Infectious Diseases 8th Edition is still the reference of choice for comprehensive global guidance on diagnosing and treating the most challenging infectious diseases Drs John E Bennett and Raphael Dolin along with new editorial team member Dr Martin Blaser have meticulously updated this latest edition to save you time and to ensure you have the latest clinical and scientific knowledge at your fingertips With new chapters expanded and updated coverage increased worldwide perspectives and many new contributors Mandell Douglas and Bennett s Principles and Practice of Infectious Diseases 8th Edition helps you identify and treat whatever infectious disease you see Get the answers to questions you have with more in depth coverage of epidemiology etiology pathology microbiology immunology and treatment of infectious agents than you ll find in any other infectious disease resource Find the latest diagnoses and treatments for currently recognized and newly emerging infectious diseases such as those caused by avian and swine influenza viruses Put the latest knowledge to work in your practice with new or completely revised chapters on influenza new pandemic strains new Middle East respiratory syndrome MERS virus probiotics antibiotics for resistant bacteria antifungal drugs new antivirals for hepatitis B and C Clostridium difficile treatment sepsis advances in HIV prevention and treatment viral gastroenteritis Lyme disease Helicobacter pylori malaria infections in immunocompromised hosts immunization new vaccines and new recommendations and microbiome Benefit from fresh perspectives and global insights from an expanded team of international contributors Find and grasp the information you need easily and rapidly with newly added chapter summaries These bulleted templates include diagnosis therapy and prevention and are designed as a quick summary of the chapter and

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 $2 \text{ CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g}) + \text{C}_2\text{H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_4(\text{g}) + \text{O}_2(\text{g}) + \text{C}_2\text{H}_6(\text{g})$   $\Delta h = 94.5 \text{ kJ}$   
 $\text{H}_2\text{O}(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l})$   $\Delta h = 71.2 \text{ kJ}$   
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{g})$   $\Delta h = 283 \text{ kJ}$

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 $\text{C}_2\text{H}_4(\text{g}) + 3 \text{O}_2(\text{g}) \rightarrow 2 \text{CO}_2(\text{g}) + 2 \text{H}_2\text{O}(\text{l})$   $\Delta h = 1411 \text{ kJ}$  mole  
 $\text{C}_2\text{H}_6(\text{g}) + 7 \text{O}_2(\text{g}) \rightarrow 2 \text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{l})$   $\Delta h = 1560 \text{ kJ}$  mole  
 $\text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l})$   $\Delta h = 285 \text{ kJ}$  mole

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 hess s law example 12.1 determine  $\Delta h$  for the target reaction  
 $\text{C}_2\text{H}_4(\text{g}) + \text{O}_2(\text{g}) \rightarrow \text{C}_2\text{H}_6(\text{g})$   
 given the following information

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 $\text{Mg}(\text{s}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{MgO}(\text{s})$  we used

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 cohen 2016 the equations for the reactions in the experiment done are as follows  
 $\text{NaOH}(\text{s}) + \text{Na}(\text{aq}) + \text{OH}(\text{aq})$

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web mar 13 2023 be able to use hess s law to calculate reaction enthalpies first law of thermodynamics energy is the ability to do work or transfer heat work is the transfer of energy from one body to another in a sense work is energy in the process of transfer

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web thus for reaction 9 2 in which 1 mole of  $\text{H}_2\text{SO}_4$  and 1 mole of  $\text{OH}^-$  combine to form 2 moles of  $\text{H}_2\text{O}$   $\Delta h = 55.8 \text{ kJ}$  as shown note that the enthalpy change  $\Delta h_{\text{products}} - \Delta h_{\text{reactants}}$  is positive if heat is absorbed that is if  $\Delta h_{\text{products}} > \Delta h_{\text{reactants}}$

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web objective the purpose of the lab is to combine two reactants in the calorimeter and measure the heats of reaction in

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