

Magnitude of the Equilibrium Constant

For a reaction



the equilibrium constant K is defined as the ratio of products to reactants at equilibrium:

$$K = \frac{[C]^c [D]^d}{[A]^a [B]^b}$$

Because K represents the ratio of products to reactants, the magnitude of K is an indicator of the extent of products and reactants present when the reaction is at equilibrium. Some chemical reactions proceed almost fully to products, whereas other chemical reactions hardly react with each other at all, or react so as to have significant amounts of both reactants and products at equilibrium.

Part A

Hydrogen is used as a rocket fuel because it is very light and reacts explosively and completely with oxygen. For the combustion reaction

$$2H_2(g) + O_2(g) \rightleftharpoons 2H_2O(g)$$

what is the likely magnitude of the equilibrium constant K ?

• **Hint**

Hint 1. How to approach the problem: Look to see if

- ☐ $K < 10^{-4}$
- ☐ $10^{-4} < K < 10^2$
- ☐ $K > 1$
- ☐ $K > 10^2$

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If the were true, the product would be formed in the reaction.

Part B

The industrial production of lime (CaO) from natural carbonates is accomplished via the following reaction:



Given the following data:

Temperature (K)	K

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