



balance nuclear reaction equations

aluminum-26 undergoes positron emission
potassium-40 undergoes beta decay (electron)
uranium-238 undergoes alpha decay

electron configurations

Write the electron configuration for C.

Write the electron configuration for Cr.

Write the shorthand notation for Bi.



NOMENCLATURE (naming compounds, writing formulas)

Write the formula for each of the following compounds:

a. Sulfur hexafluoride SF_6

b. Lithium nitride Li_3N

c. Chromium (III) carbonate $\text{Cr}_2(\text{CO}_3)_3$

d. Tin (II) chloride SnCl_2

e. Ammonium acetate $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$ (or $\text{CH}_3\text{COONH}_4$)

f. Mercury (I) chloride Hg_2Cl_2

g. Potassium bromate KBrO_3

h. Hydrosulfuric acid $\text{H}_2\text{S(OH)}_2$

i. Chloric acid HClO_3

j. Sulfurous acid H_2SO_3

Name each of the following compounds.

a. CuSO_4 Copper (II) sulfate

b. AlF_3 aluminum fluoride

c. HI hydrogen iodide

d. NO nitrogen monoxide

e. H_2Se hydrogen selenide

f. HNO_3 nitric acid

g. HNO_2 nitrous acid

h. NaHSO_4 sodium hydrogen sulfate

Half-life

Tritium (H-3) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it takes for a 40.0 g sample to decay down to 1.25 g?

$$40 \xrightarrow{1} 20 \xrightarrow{2} 10 \xrightarrow{3} 5 \xrightarrow{4} 2.5 \xrightarrow{5} 1.25$$

Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?

$$3 \text{ half lives} \quad 100 \xrightarrow{1} 50 \xrightarrow{2} 25 \xrightarrow{3} 12.5 \text{ mg}$$

Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.

$$\begin{array}{r} 63 \times .692 = 43.596 \rightarrow 43 \\ 65 \times .308 = 20.02 \rightarrow 20 \\ \hline 63.62 \rightarrow 64 \end{array}$$

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1. Understanding the eBook Pearson Education Answer Key Chemistry Study Guide
 - The Rise of Digital Reading Pearson Education Answer Key Chemistry Study Guide
 - Advantages of eBooks Over Traditional Books
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 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pearson Education Answer Key Chemistry Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pearson Education Answer Key Chemistry Study Guide
 - Personalized Recommendations

- Pearson Education Answer Key Chemistry Study Guide User Reviews and Ratings
- Pearson Education Answer Key Chemistry Study Guide and Bestseller Lists

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- Participating in Virtual Book Clubs
- Following Authors and Publishers Pearson Education Answer Key Chemistry Study Guide

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- Creating a Diverse Reading Collection Pearson Education Answer Key Chemistry Study Guide

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- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

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- Setting Reading Goals Pearson Education Answer Key Chemistry Study Guide
- Carving Out Dedicated Reading Time

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- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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