



balance nuclear reaction equations

aluminum-26 undergoes positron emission
potassium-40 undergoes beta decay (electron)
uranium-238 undergoes alpha decay

electron configurations

Write the electron configuration for C.

Write the electron configuration for Cr.

Write the shorthand notation for Bi.



NOMENCLATURE (naming compounds, writing formulas)

Write the formula for each of the following compounds:

- a. Sulfur hexafluoride SF_6
- b. Lithium nitride Li_3N
- c. Chromium (III) carbonate $\text{Cr}_2(\text{CO}_3)_3$
- d. Tin (II) chloride SnCl_2
- e. Ammonium acetate $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$ (or $\text{CH}_3\text{COONH}_4$)
- f. Mercury (I) chloride Hg_2Cl_2
- g. Potassium bromate KBrO_3
- h. Hydrosulfuric acid $\text{H}_2\text{S}(\text{aq})$
- i. Chloric acid HClO_3
- j. Sulfurous acid H_2SO_3

Name each of the following compounds.

- a. CuSO_4 Copper (II) sulfate
- b. AlF_3 aluminum fluoride
- c. HI hydrogen iodide
- d. NO nitrogen monoxide
- e. H_2Se hydroseconic acid
- f. HNO_3 nitric acid
- g. HNO_2 nitrous acid
- h. NaHSO_4 sodium hydrogen sulfate

Half-life

Tritium (H-3) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it take for a 40.0 g sample to decay down to 1.25 g?

$$40 \xrightarrow{1} 20 \xrightarrow{2} 10 \xrightarrow{3} 5 \xrightarrow{4} 2.5 \xrightarrow{5} 1.25$$

Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?

$$3 \text{ half lives} \quad 100 \xrightarrow{1} 50 \xrightarrow{2} 25 \xrightarrow{3} 12.5 \text{ mg}$$

Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.

$$\begin{array}{r} 63 \times .692 = 43.596 \rightarrow 43 \\ 65 \times .308 = 20.02 \rightarrow 20 \\ \hline 63.62 \rightarrow 64 \end{array}$$

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Robert W. Messler, Jr.

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 - The Rise of Digital Reading Pearson Chemistry Guided Practice Problem Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Pearson Chemistry Guided Practice Problem Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pearson Chemistry Guided Practice Problem Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pearson Chemistry Guided Practice Problem Answers
 - Personalized Recommendations
 - Pearson Chemistry Guided Practice Problem Answers User Reviews and Ratings
 - Pearson Chemistry Guided Practice Problem Answers and Bestseller Lists
5. Accessing Pearson Chemistry Guided Practice Problem Answers Free and Paid eBooks
 - Pearson Chemistry Guided Practice Problem Answers Public Domain eBooks
 - Pearson Chemistry Guided Practice Problem Answers eBook Subscription Services
 - Pearson Chemistry Guided Practice Problem Answers Budget-Friendly Options
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- Pearson Chemistry Guided Practice Problem Answers Compatibility with Devices
- Pearson Chemistry Guided Practice Problem Answers Enhanced eBook Features

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- Adjustable Fonts and Text Sizes of Pearson Chemistry Guided Practice Problem Answers
- Highlighting and Note-Taking Pearson Chemistry Guided Practice Problem Answers
- Interactive Elements Pearson Chemistry Guided Practice Problem Answers

8. Staying Engaged with Pearson Chemistry Guided Practice Problem Answers

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- Participating in Virtual Book Clubs
- Following Authors and Publishers Pearson Chemistry Guided Practice Problem Answers

9. Balancing eBooks and Physical Books Pearson Chemistry Guided Practice Problem Answers

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Pearson Chemistry Guided Practice Problem Answers

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Pearson Chemistry Guided Practice Problem Answers

- Setting Reading Goals Pearson Chemistry Guided Practice Problem Answers
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Pearson Chemistry Guided Practice Problem Answers

- Fact-Checking eBook Content of Pearson Chemistry Guided Practice Problem Answers
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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