

**16****SOLUTIONS****SECTION 16.1 PROPERTIES OF SOLUTIONS (pages 471–477)**

This section identifies the factors that affect the solubility of a substance and determine the rate at which a solute dissolves.

**► Solution Formation (pages 471–472)**

Look at Figure 16.1 on page 471 to help you answer Questions 1 and 2.

1. Underline the condition that causes sugar to dissolve faster in water.
  - a. as a whole cube      or      in granulated form?
  - b. when allowed to stand      or      when stirred?
  - c. at a higher temperature      or      a lower temperature?
2. Name three factors that influence the rate at which a solute dissolves in a solvent.
  - a. agitation (stirring or shaking)
  - b. temperature
  - c. particle size
3. Is the following sentence true or false? Finely ground particles dissolve more rapidly than larger particles because finer particles expose a greater surface area to the colliding solvent molecules. True

**► Solubility (pages 472–473)**

4. Complete the following table showing the steps in a procedure to determine the total amount of sodium chloride that will dissolve in 100 g of water at 25°C.

Procedure	Amount Dissolved	Amount Not Dissolved
Add 36.0 g of sodium chloride to the water	36.0 g	0.0 g
Add an additional 1.0 g of sodium chloride	0.2 g	0.6 g
Determine the total amount that dissolves	36.2 g	

5. The amount of a substance that dissolves in a given quantity of solvent at a constant temperature is called the substance's solubility at that temperature.
6. If a solution contains the maximum amount of solute for a given quantity of solvent at a constant temperature, it is called a(n) saturated solution.

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**Arshad Iqbal**

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*British Book News*, 1990

**Children's Books in Print**, 1993

[Medical Books and Serials in Print](#) R. R. Bowker LLC, R.R. Bowker Company, 1978

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