

1. States of matter

SC1.1w States of matter

- Solid:** particles are close together; particles are regularly arranged; particles can only vibrate about fixed positions; diagram chosen matches the written descriptions.
Liquid: particles are close together; particles are randomly arranged; particles can move around each other; diagram chosen matches the written descriptions.
Gas: particles are far apart; particles are randomly arranged; particles can move quickly in all directions; diagram chosen matches the written descriptions.
- Student compares the arrangements in each state – regularly arranged in solids; randomly arranged in gases and liquids. Answer compares the movement in each state – can only vibrate in solids; can move around each other in liquids; can move quickly in all directions in gases

SC1.3w Changing states – Strengthen

- Suitable particle diagrams drawn in three boxes. **Gas:** particles separate; (at least 10 diameters apart if possible); random arrangement. **Liquid:** particles touching; random arrangement. **Solid:** particles touching; regular arrangement
- Descriptions of arrangement and movements of particles in each state, written next to relevant particle diagrams. **Gas arrangement:** random; far apart. **Gas movement:** rapid in all directions. **Liquid arrangement:** random; close together. **Liquid movement:** can move around each other. **Solid arrangement:** regular; close together. **Solid movement:** can only vibrate about fixed positions
- Correct state change names written in three boxes; with descriptions of changes to arrangement and movement of particles included.

Energy at top: high; energy at bottom: low
- Should include the idea that the temperature will rise above 0 °C (melting point of water) in the spring or summer; so the ice hotel will melt.

SC1.4w Changing states – Extend

- A** Energy is transferred from the surroundings, to the particles.
B Energy is transferred to the surroundings, from the particles

- A** **Melting:** become able to move around each other; regular to random arrangement. **Freezing:** can no longer move around each other; random to regular arrangement
B **Boiling:** still randomly arranged; but become far apart; and free to move in all directions. **Condensing:** still randomly arranged; but become close together; and can only move around each other
- A** Particles become close together; and regularly arranged (rather than randomly arranged); can no longer move in all directions; but vibrate in fixed positions
B Particles become far apart; and randomly arranged; rather than regularly arranged; can move in all directions; rather than just vibrate in fixed positions
- A** The attractive forces must be weak; if they are easily overcome / formed
B It is cold and solid; it is dry because it does not melt to form a liquid

SC1.5w Particle theory and states of matter – Homework 1

- a** Liquid
b Solid
c Gas
- Particles are still randomly arranged; but become far apart and free to move in all directions
- No new substances are formed
- a** Horizontal part of curve labelled
b 55 °C labelled
c Curve from 0–7 minutes labelled
d Curve at 0 minutes labelled
- Solid at –150 °C; gas at 25 °C
- Its melting point must be below 25 °C; its boiling point must be above 25 °C.

SC1.6w Particle theory and states of matter – Homework 2

- a** Solid; liquid; gas
b **Melting:** solid → liquid; **sublimation:** solid → gas; **boiling/evaporation:** liquid → gas; **condensing:** gas → liquid; **deposition:** gas → solid; **freezing/solidifying:** liquid → solid
c Solid
- a** Close together; randomly arranged; can move around each other
b Close together; regularly arranged; can only vibrate about fixed positions

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