

1. States of matter

SC1.1w States of matter

- 1** Solid: particles are close together; particles are regularly arranged; particles can only vibrate about fixed positions; diagram chosen matches the written descriptions.

Liquid: particles are close together; particles are randomly arranged; particles can move around each other; diagram chosen matches the written descriptions.

Gas: particles are far apart; particles are randomly arranged; particles can move quickly in all directions; diagram chosen matches the written descriptions.
- 2** Student compares the arrangements in each state – regularly arranged in solids; randomly arranged in gases and liquids. Answer compares the movement in each state – can only vibrate in solids; can move around each other in liquids; can move quickly in all directions in gases.

SC1.3w Changing states – Strengthen

- 1** Suitable particle diagrams drawn in three boxes. Gas: particles separate; (at least 10 diameters apart if possible); random arrangement. Liquid: particles touching; random arrangement. Solid: particles touching; regular arrangement.
- 2** Descriptions of arrangement and movements of particles in each state, written next to relevant particle diagrams. Gas arrangement: random; far apart. Gas movement: rapid in all directions. Liquid arrangement: random; close together. Liquid movement: can move around each other. Solid arrangement: regular; close together. Solid movement: can only vibrate about fixed positions.
- 3** Correct state change names written in three boxes; with descriptions of changes to arrangement and movement of particles included.

Energy at top: high; energy at bottom: low
- 4** Should include the idea that the temperature will rise above 0 °C (melting point of water) in the spring or summer; so the ice hotel will melt.

SC1.4w Changing states – Extend

- 1** **A** Energy is transferred from the surroundings, to the particles

B Energy is transferred to the surroundings, from the particles

- 2** **A** Melting: become able to move around each other; regular to random arrangement. Freezing: can no longer move around each other; random to regular arrangement

B Boiling: still randomly arranged; but become far apart; and free to move in all directions. Condensing: still randomly arranged; but become close together; and can only move around each other
- 3** **A** Particles become close together; and regularly arranged (rather than randomly arranged); can no longer move in all directions; but vibrate in fixed positions

B Particles become far apart; and randomly arranged; rather than regularly arranged; can move in all directions; rather than just vibrate in fixed positions
- 4** **A** The attractive forces must be weak; if they are easily overcome / formed

B It is cold and solid; it is dry because it does not melt to form a liquid

SC1.5w Particle theory and states of matter – Homework 1

- 1** a Liquid

b Solid

c Gas
- 2** Particles are still randomly arranged; but become far apart and free to move in all directions
- 3** No new substances are formed
- 4** a Horizontal part of curve labelled

b 55 °C labelled

c Curve from 0–7 minutes labelled

d Curve at 0 minutes labelled
- 5** Solid at –150 °C; gas at 25 °C
- 6** Its melting point must be below 25 °C; its boiling point must be above 25 °C.

SC1.6w Particle theory and states of matter – Homework 2

- 1** a Solid; liquid; gas

b Melting: solid → liquid; sublimation: solid → gas; boiling/evaporation: liquid → gas; condensing: gas → liquid; deposition: gas → solid; freezing/solidifying: liquid → solid

c Solid
- 2** a Close together; randomly arranged; can move around each other

b Close together; regularly arranged; can only vibrate about fixed positions

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Table of Contents Pearson Chemistry Answer Key Ch 17

1. Understanding the eBook Pearson Chemistry Answer Key Ch 17
 - The Rise of Digital Reading Pearson Chemistry Answer Key Ch 17
 - Advantages of eBooks Over Traditional Books
2. Identifying Pearson Chemistry Answer Key Ch 17
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pearson Chemistry Answer Key Ch 17
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pearson Chemistry Answer Key Ch 17
 - Personalized Recommendations
 - Pearson Chemistry Answer Key Ch 17 User Reviews and Ratings
 - Pearson Chemistry Answer Key Ch 17 and Bestseller Lists
5. Accessing Pearson Chemistry Answer Key Ch 17 Free and Paid eBooks

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- Pearson Chemistry Answer Key Ch 17 Public Domain eBooks
 - Pearson Chemistry Answer Key Ch 17 eBook Subscription Services
 - Pearson Chemistry Answer Key Ch 17 Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
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 - Pearson Chemistry Answer Key Ch 17 Enhanced eBook Features
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 - Adjustable Fonts and Text Sizes of Pearson Chemistry Answer Key Ch 17
 - Highlighting and Note-Taking Pearson Chemistry Answer Key Ch 17
 - Interactive Elements Pearson Chemistry Answer Key Ch 17
 8. Staying Engaged with Pearson Chemistry Answer Key Ch 17
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pearson Chemistry Answer Key Ch 17
 9. Balancing eBooks and Physical Books Pearson Chemistry Answer Key Ch 17
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pearson Chemistry Answer Key Ch 17
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Pearson Chemistry Answer Key Ch 17
 - Setting Reading Goals Pearson Chemistry Answer Key Ch 17
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Pearson Chemistry Answer Key Ch 17
 - Fact-Checking eBook Content of Pearson Chemistry Answer Key Ch 17
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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