

Measurements Using Electrochemical Cells and Electroplating

The basic counting unit in chemistry, the mole, has a special name, *Avogadro's number*, in honor of the Italian scientist Amadeo Avogadro (1776-1856). The commonly accepted definition of Avogadro's number is the number of atoms in exactly 12 g of the isotope ^{12}C , and the quantity itself is $6.02214199(47) \times 10^{23}$.¹ A bit of information about Avogadro seems appropriate. His full name was Lorenzo Romano Amedeo Carlo Avogadro (almost a mole of letters in his name). He was a practicing lawyer until 1806 when he began his new career teaching physics and math at the University of Turin, where he was later promoted to the chair of physical chemistry. In 1811, Avogadro published a paper in the *Journal de Physique*, entitled "Essay on a Manner of Determining the Relative Masses of the Elementary Molecules of Bodies, and the Proportions in Which They Enter into These Compounds," which pretty much says it all. This paper includes the statement that has come to be regarded as Avogadro's Hypothesis:

The first hypothesis to present itself in this connection, and apparently even the only admissible one, is the supposition that the number of integral molecules in any gases is always the same for equal volumes, or always proportional to the volumes. Indeed, if we were to suppose that the number of molecules contained in a given volume were different for different gases, it would scarcely be possible to conceive that the law regulating the distance of molecules could give in all cases relations as simple as those which the facts just detailed compel us to acknowledge between the volumes and the number of molecules.

In this experiment, you will confirm Avogadro's number by conducting an electrochemical process called electrolysis. In electrolysis, an external power supply is used to drive an otherwise nonspontaneous reaction. You will use a copper strip and a zinc strip as the electrodes, placed in a beaker of sulfuric acid. You will make the cell electrolytic by using the copper strip as the anode and the zinc strip as the cathode. By determining the average current used in the reaction, along with the knowledge that all of the copper ions formed are the 2^+ cations, you will calculate the number of atoms in one mole of copper and compare this value with Avogadro's number.

OBJECTIVES

In this experiment, you will

- Prepare an electrochemical cell to oxidize a copper electrode.
- Measure the amount of copper that was deposited in the electroplating process and determine the average current used.
- Calculate a value for Avogadro's number and compare it to the accepted value.
- Calculate the value of a Faraday.

MATERIALS

Lab Quest

Current probe

Battery pack with 2 D batteries

2 wires with banana plug

3 wires with alligator clips

1 M H_2SO_4 solution

steel wool or sandpaper

balance

distilled water

copper strip (anode) (+)*

zinc strip (cathode) (-)*

* Remember the acronym "EPA"—An Electrolytic (nonspontaneous) cell has a Positive Anode.

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Yi-Tong Ma

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Table of Contents Measurements Using Electrochemical Cells And Electroplating

1. Understanding the eBook Measurements Using Electrochemical Cells And Electroplating
 - The Rise of Digital Reading Measurements Using Electrochemical Cells And Electroplating
 - Advantages of eBooks Over Traditional Books
2. Identifying Measurements Using Electrochemical Cells And Electroplating
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Measurements Using Electrochemical Cells And Electroplating
 - User-Friendly Interface
4. Exploring eBook Recommendations from Measurements Using Electrochemical Cells And Electroplating
 - Personalized Recommendations
 - Measurements Using Electrochemical Cells And Electroplating User Reviews and Ratings
 - Measurements Using Electrochemical Cells And Electroplating and Bestseller Lists

5. Accessing Measurements Using Electrochemical Cells And Electroplating Free and Paid eBooks
 - Measurements Using Electrochemical Cells And Electroplating Public Domain eBooks
 - Measurements Using Electrochemical Cells And Electroplating eBook Subscription Services
 - Measurements Using Electrochemical Cells And Electroplating Budget-Friendly Options
6. Navigating Measurements Using Electrochemical Cells And Electroplating eBook Formats
 - ePUB, PDF, MOBI, and More
 - Measurements Using Electrochemical Cells And Electroplating Compatibility with Devices
 - Measurements Using Electrochemical Cells And Electroplating Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Measurements Using Electrochemical Cells And Electroplating
 - Highlighting and Note-Taking Measurements Using Electrochemical Cells And Electroplating
 - Interactive Elements Measurements Using Electrochemical Cells And Electroplating
8. Staying Engaged with Measurements Using Electrochemical Cells And Electroplating
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Measurements Using Electrochemical Cells And Electroplating
9. Balancing eBooks and Physical Books Measurements Using Electrochemical Cells And Electroplating
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Measurements Using Electrochemical Cells And Electroplating
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Measurements Using Electrochemical Cells And Electroplating
 - Setting Reading Goals Measurements Using Electrochemical Cells And Electroplating
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Measurements Using Electrochemical Cells And Electroplating
 - Fact-Checking eBook Content of Measurements Using Electrochemical Cells And Electroplating
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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