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**Didier Musso**

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we can write a balanced nuclear equation for this reaction as let's do a practice example identify the product x by writing a balanced nuclear equation  $^{137}_{55}\text{Cs} + ^{137}_{56}\text{Ba} \rightarrow ^{137}_{56}\text{Ba} + \text{x}$  because the mass number is the same on both sides of the equation x must have a

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web answer 4 correct in the reactants and products must be the same part b what is the value of in the following nuclear reaction  ${}_{6}^{15}\text{I} + {}_{9}^{19}\text{F} \rightarrow {}_{6}^{15}\text{X} + {}_{1}^{1}\text{H}$  express your answer as an integer hint 1 how to approach the problem use rule 2 the total of the subscripts atomic numbers mathematically this can be written as answer

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