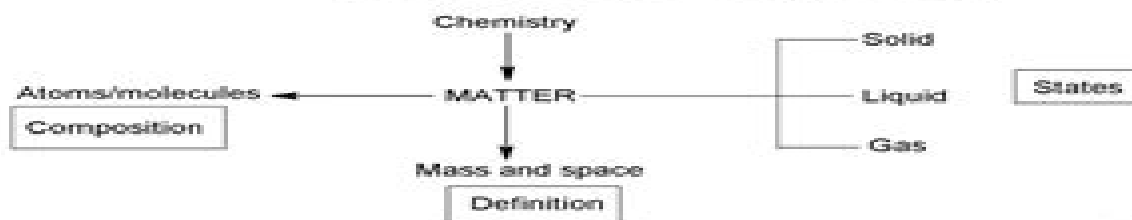


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Atoms: Building block of matter. It may or may not exist free in nature

Examples: H, He, Na, C

Shells = electrons

Mass = 0

Charge = -1



Nucleus = Protons + Neutrons

Charge = +1 0

Mass = 1 1

Nucleus: It is central part of an atom.

Shells: Circular paths around the nucleus where electrons reside. They are also called orbits or energy levels.

No. of es in a shell = $2n^2$

Where n = no of shell.

1st shell = 2es, 2nd shell = 8es, 3rd shell = 18es.

At least 8 electron should be present in last for the stability of an atom. However if an atom has one shell like Hydrogen and Helium then 2 electrons are required for complete shell and stability.

Atomic number:

Number of protons present in the nucleus of an atom of an element.

Examples: $_{11}\text{Na}$, $_6\text{C}$, $_8\text{O}$

Atomic mass or Mass number or nucleon number: It is the sum of protons and neutrons present in the nucleus of an atom of an element.

Examples: ^{23}Na , ^{12}C , ^{16}O

O Level Chemistry Notes

J Dewey



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